

Write Titrating Volume (Spreads: 1.00)

Name \_\_\_\_\_

Acetic Acid (pKa = 4.75) titrated with NaOH

a. 0.100 M / volume 10.00 mL / acid / base 0.100 M

b. 0.100 M / volume 10.00 mL / acid / base 0.100 M

c. 0.100 M / pH 4.75 / compound that releases acetate  
0.100 M / acid / base 0.100 M / acid

Acetic Acid:  $\text{CH}_3\text{COOH}$ ,  $\text{CH}_3\text{COO}^-$ ,  $\text{CH}_3\text{COO}^-$

$\text{CH}_3\text{COOH}$  /  $\text{CH}_3\text{COO}^-$  /  $\text{CH}_3\text{COO}^-$   
 $\text{CH}_3\text{COOH}$  /  $\text{CH}_3\text{COO}^-$  /  $\text{CH}_3\text{COO}^-$

Species	Add (M, Base (M, or both (M)	Acids with $\text{M}_1$ (M or both)	Conductivity (C or both)
$\text{CH}_3\text{COOH}$	$\text{CH}_3\text{COOH}$	$\text{CH}_3\text{COOH}$	$\text{CH}_3\text{COOH}$
$\text{CH}_3\text{COO}^-$	$\text{CH}_3\text{COO}^-$	$\text{CH}_3\text{COO}^-$	$\text{CH}_3\text{COO}^-$
$\text{NaOH}$	$\text{NaOH}$	$\text{NaOH}$	$\text{NaOH}$
$\text{NaCH}_3\text{COO}$	$\text{NaCH}_3\text{COO}$	$\text{NaCH}_3\text{COO}$	$\text{NaCH}_3\text{COO}$

# Note Taking Guide 1101 Answers

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