

CLO- Classification of Elements & Periodicity in Properties

MULTIPLE CHOICE QUESTIONS:

Choose the one correct option in the following:

- Identify the least stable ion amongst the following :
(a) Li^+ (b) Be^+ (c) B^+ (d) C^+
- The correct order of radii is :
(a) $\text{N} < \text{Be} < \text{B}$ (b) $\text{F}^- < \text{O}^{2-} < \text{N}^{3-}$ (c) $\text{N} < \text{Li} < \text{K}$ (d) $\text{Fe}^{2+} < \text{Fe}^{3+} < \text{Fe}^{4+}$
- Which one of the following has largest size :
(a) Na (b) Na^+ (c) Na^- (d) Li^+
- The number of elements in the 4th period of periodic table is :
(a) 8 (b) 10 (c) 18 (d) 32
- Which of the following has least electron affinity (electron gain enthalpy) :
(a) O (b) N (c) Ar (d) F
- An element of atomic number 29 belongs to which of the following block of the periodic table :
(a) s-block (b) p-block (c) d-block (d) f-block
- Which one of the following arrangements not truly represents the property indicated against it:
(a) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – electronegativity (b) $\text{Br}_2 < \text{F}_2 < \text{Cl}_2$ – Electron Affinity
(c) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – Bond energy (d) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – Oxidising Power
- Alkali metals in each period have :
(a) Smallest size (b) Lowest IE (c) Highest IE (d) Highest EN
- A sudden large jump between the value of second and third ionization energies of elements would be associated with which of the following electronic configuration :
(a) $2\text{p}^4 3\text{s}^1$ (b) $3\text{s}^2 3\text{p}^1$ (c) $3\text{s}^2 3\text{p}^2$ (d) $2\text{p}^4 3\text{s}^2$
- Which of the following pairs are chemically dissimilar :
(a) Na and K (b) Ba and Sr (c) Zr and Hf (d) Ca and Zn
- Collective name given to the elements with outer shell configuration $ns^2 np^6$ is :
(a) Chalkogens (b) Alkaline earths metals (c) Transition metals (d) Nobel gases
- Modern extended periodic table was given by :
(a) Bohr (b) Lothar mayer (c) Moseley (d) None alone
- The most electronegative elements is :
(a) Hydrogen (b) Fluorine (c) Chlorine (d) Oxygen
- The element having electronic configuration :
 $[\text{Kr}]4\text{d}^{10}4\text{f}^{14}5\text{s}^2 5\text{p}^6 5\text{d}^1 6\text{s}^2$ belongs to
(a) s-block (b) p-block (c) d-block (d) f-block
- An elements R forms the highest oxide R_2O_5 . R belong to :

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Table of Contents Periodicity Multiple Choice Questions

1. Understanding the eBook Periodicity Multiple Choice Questions
 - The Rise of Digital Reading Periodicity Multiple Choice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodicity Multiple Choice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodicity Multiple Choice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodicity Multiple Choice Questions

- Personalized Recommendations
- Periodicity Multiple Choice Questions User Reviews and Ratings
- Periodicity Multiple Choice Questions and Bestseller Lists
- 5. Accessing Periodicity Multiple Choice Questions Free and Paid eBooks
 - Periodicity Multiple Choice Questions Public Domain eBooks
 - Periodicity Multiple Choice Questions eBook Subscription Services
 - Periodicity Multiple Choice Questions Budget-Friendly Options
- 6. Navigating Periodicity Multiple Choice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodicity Multiple Choice Questions Compatibility with Devices
 - Periodicity Multiple Choice Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodicity Multiple Choice Questions
 - Highlighting and Note-Taking Periodicity Multiple Choice Questions
 - Interactive Elements Periodicity Multiple Choice Questions
- 8. Staying Engaged with Periodicity Multiple Choice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodicity Multiple Choice Questions
- 9. Balancing eBooks and Physical Books Periodicity Multiple Choice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodicity Multiple Choice Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodicity Multiple Choice Questions
 - Setting Reading Goals Periodicity Multiple Choice Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodicity Multiple Choice Questions

- Fact-Checking eBook Content of Periodicity Multiple Choice Questions
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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