

Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Fluorine
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (column 17), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



I am so perfect... I have a full valence shell. I don't need those other elements... I hope they just leave me alone!



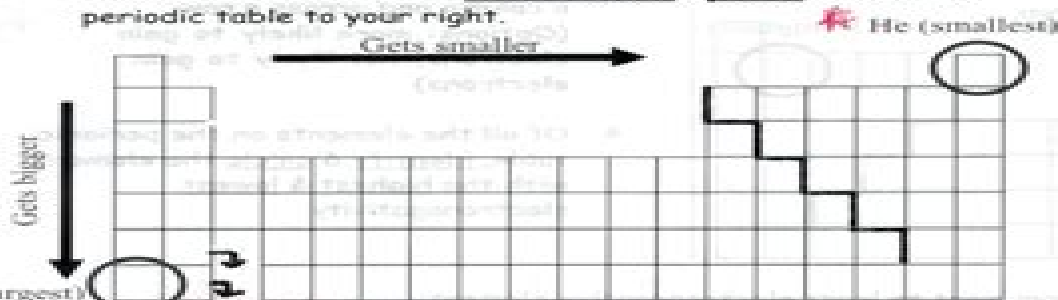
Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

Effective nuclear charge, Z_{eff} , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii

*not a good enough answer
4AP Answer*

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Table of Contents Periodic Trends Pogil Extension Questions Answer Key

1. Understanding the eBook Periodic Trends Pogil Extension Questions Answer Key
 - The Rise of Digital Reading Periodic Trends Pogil Extension Questions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Pogil Extension Questions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Pogil Extension Questions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Pogil Extension Questions Answer Key
 - Personalized Recommendations
 - Periodic Trends Pogil Extension Questions Answer Key User Reviews and Ratings
 - Periodic Trends Pogil Extension Questions Answer Key and Bestseller Lists
5. Accessing Periodic Trends Pogil Extension Questions Answer Key Free and Paid eBooks
 - Periodic Trends Pogil Extension Questions Answer Key Public Domain eBooks
 - Periodic Trends Pogil Extension Questions Answer Key eBook Subscription Services
 - Periodic Trends Pogil Extension Questions Answer Key Budget-Friendly Options
6. Navigating Periodic Trends Pogil Extension Questions Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Periodic Trends Pogil Extension Questions Answer Key Compatibility with Devices
 - Periodic Trends Pogil Extension Questions Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Pogil Extension Questions Answer Key
 - Highlighting and Note-Taking Periodic Trends Pogil Extension Questions Answer Key
 - Interactive Elements Periodic Trends Pogil Extension Questions Answer Key
 8. Staying Engaged with Periodic Trends Pogil Extension Questions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Pogil Extension Questions Answer Key
 9. Balancing eBooks and Physical Books Periodic Trends Pogil Extension Questions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Pogil Extension Questions Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Periodic Trends Pogil Extension Questions Answer Key
 - Setting Reading Goals Periodic Trends Pogil Extension Questions Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Periodic Trends Pogil Extension Questions Answer Key
 - Fact-Checking eBook Content of Periodic Trends Pogil Extension Questions Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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