

# Annotated periodic table of the elements

1		2		Group name*																13		14		15		16		17		18																											
1 H 1.01 Hydrogen				1 Hydrogen & the alkali metals																5 B 10.81 Boron		6 C 12.01 Carbon		7 N 14.01 Nitrogen		8 O 16.00 Oxygen		9 F 18.99 Fluorine		10 Ne 20.18 Neon																											
3 Li 6.94 Lithium		4 Be 9.01 Beryllium		2 Alkaline earth metals																11 Na 22.99 Sodium		12 Mg 24.31 Magnesium		13 Al 26.98 Aluminum		14 Si 28.09 Silicon		15 P 30.97 Phosphorus		16 S 32.06 Sulfur		17 Cl 35.45 Chlorine		18 Ar 39.95 Argon																							
11 Na 22.99 Sodium		12 Mg 24.31 Magnesium		3 Coinage metals																19 K 39.10 Potassium		20 Ca 40.08 Calcium		21 Sc 44.96 Scandium		22 Ti 47.87 Titanium		23 V 50.94 Vanadium		24 Cr 52.00 Chromium		25 Mn 54.94 Manganese		26 Fe 55.85 Iron		27 Co 58.93 Cobalt		28 Ni 58.69 Nickel		29 Cu 63.55 Copper		30 Zn 65.38 Zinc		31 Ga 69.72 Gallium		32 Ge 72.63 Germanium		33 As 74.92 Arsenic		34 Se 78.96 Selenium		35 Br 79.90 Bromine		36 Kr 83.80 Krypton			
19 K 39.10 Potassium		20 Ca 40.08 Calcium		4 Coinage metals																37 Rb 85.47 Rubidium		38 Sr 87.62 Strontium		39 Y 88.91 Yttrium		40 Zr 91.22 Zirconium		41 Nb 92.91 Niobium		42 Mo 95.94 Molybdenum		43 Tc 98.91 Technetium		44 Ru 101.07 Ruthenium		45 Rh 102.91 Rhodium		46 Pd 106.42 Palladium		47 Ag 107.87 Silver		48 Cd 112.41 Cadmium		49 In 114.82 Indium		50 Sn 118.71 Tin		51 Sb 121.76 Antimony		52 Te 127.60 Tellurium		53 I 126.91 Iodine		54 Xe 131.29 Xenon			
37 Rb 85.47 Rubidium		38 Sr 87.62 Strontium		5 Coinage metals																55 Cs 132.91 Cesium		56 Ba 137.33 Barium		57 to 70		71 La 138.91 Lanthanum		72 Hf 178.49 Hafnium		73 Ta 180.95 Tantalum		74 W 183.85 Tungsten		75 Re 186.21 Rhenium		76 Os 190.23 Osmium		77 Ir 192.22 Iridium		78 Pt 195.08 Platinum		79 Au 196.97 Gold		80 Hg 200.59 Mercury		81 Tl 204.38 Thallium		82 Pb 207.2 Lead		83 Bi 208.98 Bismuth		84 Po 209 Polonium		85 At 210 Astatine		86 Rn 222 Radon	
55 Cs 132.91 Cesium		56 Ba 137.33 Barium		6 Coinage metals																87 to 102		103 Fr 223 Francium		104 Ra 226 Radium		105 to 118		119 Nh 289 Nihonium		120 Fl 289 Flerovium		121 Mc 288 Moscovium		122 Lv 293 Livermorium		123 Ts 294 Tennessine		124 Og 294 Oganesson																			
87 Fr 223 Francium		88 Ra 226 Radium		7 Coinage metals																																																					

CLASS COUNTERPARTS		
Metals		Nonmetals
Active	Pre-transition Alkali earth Alkaloids	Cometoids
Transition		Intermediate
Frontier		Metalloids
Node (Ru-Pt-Co-Au)		Node gas

57 La 138.91 Lanthanum	58 Ce 140.12 Cerium	59 Pr 140.91 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 144.91 Promethium	62 Sm 150.36 Samarium	63 Eu 151.96 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.93 Terbium	66 Dy 162.50 Dysprosium	67 Ho 164.93 Holmium	68 Er 167.26 Erbium	69 Tm 168.93 Thulium	70 Yb 173.05 Ytterbium
89 Ac 227.03 Actinium	90 Th 232.04 Thorium	91 Pa 231.04 Protactinium	92 U 238.03 Uranium	93 Np 237.05 Neptunium	94 Pu 244.06 Plutonium	95 Am 243.06 Americium	96 Cm 247.07 Curium	97 Bk 247.07 Berkelium	98 Cf 251.08 Californium	99 Es 252.08 Einsteinium	100 Fm 257.10 Fermium	101 Md 258.10 Mendelevium	102 No 259.10 Nobelium

**Active metals**

**Pre-transition**  
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

**Rare earths**  
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

**Actinides**  
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

**Transition metals**  
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

**Noble metals**  
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.

**Frontier metals**  
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

**Metalloids**  
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

**Intermediate nonmetals**  
H and N are colourless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H<sub>2</sub>O, CO, NO, N<sub>2</sub>O).

**Group 1**  
This group is Sc-Y, and the alkaline earths are La-Be, and Al-Li (Be is not a metal). These group 1-2 elements are the most reactive in ambient conditions.

**Cometoids**  
O, F, and Cl are, respectively, colourless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

**Noble gases**  
Colourless, odourless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF<sub>6</sub>, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

# Periodic Table Guide Answer Sheet

**Rachel S Tattersall**



## **Periodic Table Guide Answer Sheet:**

## **Periodic Table Guide Answer Sheet** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Periodic Table Guide Answer Sheet**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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