Name: Brittany French Date: October 4, 2012

Student Exploration: Electron Configuration

| Activity B: | Get the Gizmo ready: | A STATE OF THE PARTY OF THE PAR |
|--------------|--|--|
| Atomic radii | On the PERIODIC TABLE tab, select Na (sodium). Select the ELECTRON CONFIGURATION tab. | * |

Question: How do the radii of atoms change across a period of the periodic table?

 Predict: Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the atomic radii will change as electrons are added to a shell?

It will decrease

Arrange: Create a proper electron configuration for sodium. After clicking Check, note the Electron configuration and the Atomic radius now listed at right.

Sodium electron configuration: 1s2 2s2 2p6 3s1

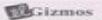
Atomic radius: 190 Pico meters

 Compare: Click Next element, and then add an electron to the magnesium atom. Click check, and record the electron configuration and atomic radius below.

Magnesium electron configuration: 1s2 2s2 2p6 3s2 Atomic radius: 145 Pico meters

 Gather data: Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is pm.)

| Element | Number of electrons | Electron configuration | Atomic radius (pm) |
|------------|------------------------|------------------------|--------------------|
| Aluminum | 13 | 1s2 2s2 2p6 3s2 3p1 | 118 |
| Silicon | 14 | 1s2 2s2 2p6 3s2 3p2 | 111 |
| Phosphorus | 15 | 1s2 2s2 2p6 3s2 3p3 | 98 |
| Sulfur | 16 | 1s2 2s2 2p6 3s2 3p4 | 88 |
| Chlorine | 17 | 1s2 2s2 2p6 3s2 3p5 | 79 |
| Argon | 18 | 1s2 2s2 2p6 3s2 3p6 | 71 |
| | | | |



Periodic Table Exploration Answer Key

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