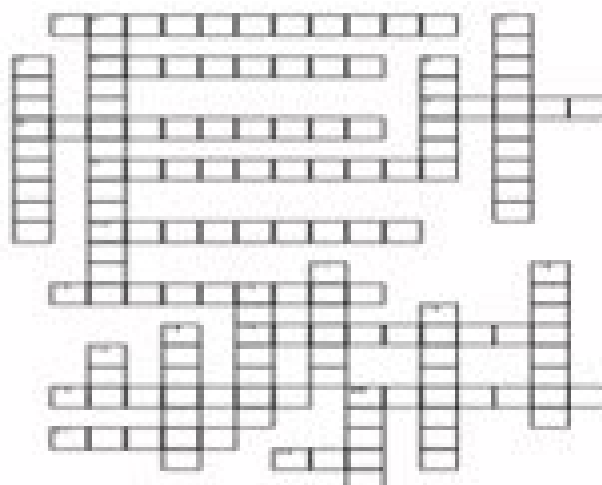


Periodic Table Assignment

Read pages 14 – 19. Complete the crossword puzzle (in pencil) and fill in the period table as instructed below.



Abstract

1. Synthesis elements with atomic numbers greater than 93
2. This law states that the properties of elements repeat as a function of their atomic number
3. Lanthan is in this period
4. Elements were originally placed in order of atomic ____, now they are in order of atomic ____
5. The group that consists of very unreactive gases
6. Number of elements known in 1869
7. Element 101 was named in his honour
8. The periodic table is normally drawn with 18 columns. A more concise representation has ____ columns
9. Most nonmetals are in this state
10. Nonmetal that is a liquid at SATP
11. The general name given to a column in the periodic table
12. According to figure 6, when aluminum combines with oxygen, three oxygen atoms combine with this many aluminum atoms.

1999

2. Elements that most closely follow the periodic law
3. The first scientist to organize elements based on repeating patterns
4. The modern name for alkali
6. Most elements are _____ since they fall to the left of the "staircase line"
11. The general name given to a row (left-to-right) in the periodic table
12. Iodine belongs to this group
14. The first alkali metal
15. This element was assigned a relative atomic mass of one because it is the lightest element
16. The state of matter represented by (aq)
18. Silicon is in this group
20. The staircase that separates metals and nonmetals begins below this element

[illegible]

- 1) Place a large G in boxes that house elements that are gases at S.A.T.P., and a large L in boxes that contain elements that are liquids. The rest of the elements are solids; do not write anything in these boxes.
- 2) Darken the "staircase line" that divides metals from non-metals.
- 3) Metalloids (i.e. semi-conductors) are elements that have properties midway between metals and non-metals. The metalloids are: B, Si, Ge, As, Sb, Te, Po, At. Identify the metalloids by drawing diagonal lines in these boxes.
- 4) Notice that each group (column) in the periodic table is identified at the top with a number and/or letter. There are two conventions used in the textbook. Label the groups that end in A (e.g. 1A, 2A, 3A, etc.).
- 5) Frame and colour the following groups (note: H is not an 1A metal): Alkali metals, alkaline earth metals, halogens, noble gases, transition metals, the inner transition elements (made up of lanthanides and actinides).
- 6) Create a legend for your periodic table.

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