

# Percent Yield Practice I

1. A student uses stoichiometry to calculate the predicted yield of  $\text{CO}_2$  in a chemical reaction as 323 g. When the student *mixes the chemicals* to make the  $\text{CO}_2$ , he finds that the reaction only produces *308 g* of  $\text{CO}_2$ . What is the percent yield?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{Percent yield} = \frac{308}{323} \times 100$$

$$\text{Percent yield} = 95.4\%$$

# Percent Yield Calculations Worksheet Key

**Tobias Bleicker**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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