

**Periodic Trends**

1. Explain the following trends:

- Atomic radius
- Electronegativity
- Ionization energy

2. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

3. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

4. An element has the following properties:

- Atomic number: 55
- Group: 1
- Period: 5

5. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

6. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

7. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

8. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

9. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

10. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

**Periodic Trends Test Study Guide**

**Atomic Structure and Properties of the Periodic Table**

1. What are the three subatomic particles found in the periodic table?

- Protons: positive charge, located in the nucleus
- Neutrons: neutral charge, located in the nucleus
- Electrons: negative charge, located outside the nucleus

2. Label the following parts of the periodic table: s-block, p-block, d-block, f-block, transition metals, main group elements, noble gases, lanthanides, actinides.

3. Complete the following table:

Block	Groups	Periods	Elements
s-block	1, 2	1-7	Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra
p-block	13-18	2-7	B, C, N, O, F, Ne, Al, Si, P, S, Cl, Ar, As, Se, Br, Kr, Sb, Te, I, Xe, Bi, Po, At, Rn
d-block	3-10	4-7	Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Ga, Ge, As, Se, Br, Kr, Rb, Sr, Y, Zr, Nb, Mo, Tc, Ru, Rh, Pd, Ag, Cd, In, Sn, Sb, Te, I, Xe, Ba, La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu, Hf, Ta, W, Re, Os, Ir, Pt, Au, Hg, Tl, Pb, Bi, Po, At, Rn, Ac, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr
f-block	14-15	6-7	Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr

4. Explain the periodic trends of atomic radius, electronegativity, ionization energy, and electron affinity.

5. Complete the following table:

Property	Increases from	Decreases from
Atomic radius	Top-left	Bottom-right
Electronegativity	Bottom-right	Top-left
Ionization energy	Top-left	Bottom-right
Electron affinity	Top-right	Bottom-left

6. Explain the periodic trends of atomic radius, electronegativity, ionization energy, and electron affinity.

7. Complete the following table:

Property	Increases from	Decreases from
Atomic radius	Top-left	Bottom-right
Electronegativity	Bottom-right	Top-left
Ionization energy	Top-left	Bottom-right
Electron affinity	Top-right	Bottom-left

8. Explain the periodic trends of atomic radius, electronegativity, ionization energy, and electron affinity.

9. Complete the following table:

Property	Increases from	Decreases from
Atomic radius	Top-left	Bottom-right
Electronegativity	Bottom-right	Top-left
Ionization energy	Top-left	Bottom-right
Electron affinity	Top-right	Bottom-left

10. Explain the periodic trends of atomic radius, electronegativity, ionization energy, and electron affinity.

11. Complete the following table:

Property	Increases from	Decreases from
Atomic radius	Top-left	Bottom-right
Electronegativity	Bottom-right	Top-left
Ionization energy	Top-left	Bottom-right
Electron affinity	Top-right	Bottom-left

# Periodic Table Study Guide Word Search

**Chao Zhang**



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## **Table of Contents Periodic Table Study Guide Word Search**

1. Understanding the eBook Periodic Table Study Guide Word Search
  - The Rise of Digital Reading Periodic Table Study Guide Word Search
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Study Guide Word Search
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Study Guide Word Search
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Study Guide Word Search
  - Personalized Recommendations
  - Periodic Table Study Guide Word Search User Reviews and Ratings

- Periodic Table Study Guide Word Search and Bestseller Lists
- 5. Accessing Periodic Table Study Guide Word Search Free and Paid eBooks
  - Periodic Table Study Guide Word Search Public Domain eBooks
  - Periodic Table Study Guide Word Search eBook Subscription Services
  - Periodic Table Study Guide Word Search Budget-Friendly Options
- 6. Navigating Periodic Table Study Guide Word Search eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Table Study Guide Word Search Compatibility with Devices
  - Periodic Table Study Guide Word Search Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Table Study Guide Word Search
  - Highlighting and Note-Taking Periodic Table Study Guide Word Search
  - Interactive Elements Periodic Table Study Guide Word Search
- 8. Staying Engaged with Periodic Table Study Guide Word Search
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Table Study Guide Word Search
- 9. Balancing eBooks and Physical Books Periodic Table Study Guide Word Search
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Table Study Guide Word Search
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Study Guide Word Search
  - Setting Reading Goals Periodic Table Study Guide Word Search
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Study Guide Word Search
  - Fact-Checking eBook Content of Periodic Table Study Guide Word Search
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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