

CECIL CONVENT SCHOOL

PERIODIC ASSESSMENT-2

Class 11 - Chemistry

Time Allowed: 1 hour

Maximum Marks: 20

General Instructions:

1. There are 11 questions in this question paper.
2. SECTION A - Q. No. 1 to 5 are very short answer questions carrying 1 mark each.
3. SECTION B - Q. No. 6 to 8 are short answer questions carrying 2 marks each.
4. SECTION C- Q. No. 9 to 11 is case based question carrying 3 marks each.
5. All questions are compulsory.
6. Attempt questions in order.

1. Water can be lifted into the water tank at the top of the house with the help of a pump. Then why is it not considered to be spontaneous? [1]
2. Neither q nor W is a state function but $q + W$ is a state function. Explain why? [1]
3. What is the enthalpy change for an adiabatic process? [1]
4. State the law of chemical equilibrium. [1]
5. A tank is full of water. Water is coming in as well as going out at same rate. What will happen to level of water in a tank? What is name given to such state? [1]
6. Calculate the standard enthalpy of formation of $\text{CH}_3\text{OH}(\text{l})$ from the following data: [2]



7. The equilibrium constant for a reaction is 10. What will be the value of ΔG^\ominus ? $R = 8.314 \text{ JK}^{-1} \text{mol}^{-1}$, $T = 300 \text{ K}$. [2]
8. Write the equilibrium constant (K_c) expression for the following reactions. [2]



9. i. Why standard entropy of an elementary substance is not zero whereas standard enthalpy of formation is taken as zero? [3]
 ii. Under what conditions will the reaction occur, if
 a. both ΔH and ΔS are positive
 b. both ΔH and ΔS are negative
10. Dihydrogen gas is obtained from the natural gas by partial oxidation with steam as per following endothermic reaction: [1]
 $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$
 a. Write the expression for K_p for the above reaction.

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