

Step 8. Periodic table trends

Electronegativity

Ionization energy

He

Atomic radius

The periodic table displays various elements with their atomic numbers and symbols. Three main trends are highlighted with arrows:

- Atomic radius:** Indicated by a red arrow pointing downwards on the left side of the table.
- Ionization energy:** Indicated by a green arrow pointing upwards on the right side of the table.
- Electronegativity:** Indicated by a black arrow pointing upwards on the right side of the table.

Two diagonal arrows across the table represent other trends:

- Non-metallic properties:** Indicated by a purple arrow pointing from the bottom-left towards the top-right.
- Metallic properties:** Indicated by a blue arrow pointing from the top-left towards the bottom-right.

Atomic radius

An arrow head shows the INCREASE of the property

Example: comparing K atom with Br atom

- Atomic radius of K is HIGHER than atomic radius of Br
- Ionization energy of K is LOWER than ionisation energy of Br
- Metallic properties of K are HIGHER than metallic properties of Br



Periodic Trends Key

Anthony S. Fauci



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