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Ionization energy



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Atomic radius

An arrow head shows the INCREASE of the property

Example: comparing K atom with Br atom

- . Atomic radius of K is HIGHER than atomic radius of Br
- · Ionization energy of K is LOWER than ionisation energy of Br
- Metallic properties of K are HIGHER than metallic properties of Br

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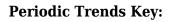
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Periodic Trends Key

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