Name: Key

Date:

Honors Chemistry

#### Percent Yield Practice

 Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?

2 Al + 3 CuSO<sub>4</sub> - 3 Cu + Al<sub>2</sub>(SO<sub>4</sub>);

1. Start . 1 met Al . 3 met CuSO<sub>4</sub> 159.62 g (uSO<sub>4</sub> = 16.6 g CuSO<sub>4</sub>.

26.98 g Al . 2 met Al . 1 met CuSO<sub>4</sub> = 16.6 g CuSO<sub>4</sub>.

9.65 g CuSO<sub>4</sub> . 1 not CusO<sub>4</sub> . 3 met Cu . 63.55 g Cu = 3.84 g Cu

159.62 g CuSO<sub>4</sub> . 3 met Cu . 63.55 g Cu = 3.84 g Cu

159.62 g CuSO<sub>4</sub> . 3 met Cu . 1 met CuSO<sub>4</sub> . 1 met CuSO<sub>4</sub>

3. Fyou start with 5.0 g of C<sub>3</sub>H<sub>8</sub> and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?

ZNaCI + \_ CaO - \_ CaCI2 + Na2O

ZOg CaO . | mot CaO . / not tac 61.98 g Maz 0 = 200 g Maz 0 |

56.08 g CaO . | mot CaO . | mot

# **Percent Yield Answer Key**

**Rachel S Tattersall** 

#### **Percent Yield Answer Key:**

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