

Reaction Rates and Potential Energy Diagrams

1. Chemical reactions occur when reactants collide. For what reasons may a collision fail to produce a chemical reaction?

Not enough energy; improper angle.

2. If every collision between reactants leads to a reaction, what determines the rate at which the reaction occurs?

Nature of reactants, Concentration, Temperature, Catalysts.

3. What is the activation energy of a reaction, and how is this energy related to the activated complex of the reaction?

E_a is the minimum amount of energy for a reaction to occur. It is the amount of energy required to create an activated complex.

4. What happens when a catalyst is used in a reaction?

A catalyst changes the reaction mechanism, in the process lowering the activation energy.

5. Name 4 things that will speed up or slow down a chemical reaction.

Increase concentration by distillation of a solvent, Increase concentration by increasing pressure of a gas, Increase temp, Add a catalyst, Add an inhibitor.

6. Draw an energy diagram for a reaction. Label the axis, PE of reactants = 350 KJ/mol, E_a = 100 KJ/mol, PE of products = 250 KJ/mol.



7. Is the reaction in # 6 exothermic or endothermic? Explain.

Exothermic. The ΔH is -100 KJ/mol which means heat is released.

8. How could you lower the activation energy for the reaction in #6?

Add a catalyst.

Potential Energy Diagram Practice Problems

**Patrick E. McMahon, Rosemary
McMahon, Bohdan Khomtchouk**



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Table of Contents Potential Energy Diagram Practice Problems

1. Understanding the eBook Potential Energy Diagram Practice Problems
 - The Rise of Digital Reading Potential Energy Diagram Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Potential Energy Diagram Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Potential Energy Diagram Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Potential Energy Diagram Practice Problems
 - Personalized Recommendations
 - Potential Energy Diagram Practice Problems User Reviews and Ratings
 - Potential Energy Diagram Practice Problems and Bestseller Lists
5. Accessing Potential Energy Diagram Practice Problems Free and Paid eBooks
 - Potential Energy Diagram Practice Problems Public Domain eBooks
 - Potential Energy Diagram Practice Problems eBook Subscription Services
 - Potential Energy Diagram Practice Problems Budget-Friendly Options

6. Navigating Potential Energy Diagram Practice Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Potential Energy Diagram Practice Problems Compatibility with Devices
 - Potential Energy Diagram Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Potential Energy Diagram Practice Problems
 - Highlighting and Note-Taking Potential Energy Diagram Practice Problems
 - Interactive Elements Potential Energy Diagram Practice Problems
8. Staying Engaged with Potential Energy Diagram Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Potential Energy Diagram Practice Problems
9. Balancing eBooks and Physical Books Potential Energy Diagram Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Potential Energy Diagram Practice Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Potential Energy Diagram Practice Problems
 - Setting Reading Goals Potential Energy Diagram Practice Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Potential Energy Diagram Practice Problems
 - Fact-Checking eBook Content of Potential Energy Diagram Practice Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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