

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Trends And The Properties Of Elements Lab Answers

Tobias Bleicker



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Table of Contents Periodic Trends And The Properties Of Elements Lab Answers

1. Understanding the eBook Periodic Trends And The Properties Of Elements Lab Answers
 - The Rise of Digital Reading Periodic Trends And The Properties Of Elements Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends And The Properties Of Elements Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends And The Properties Of Elements Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends And The Properties Of Elements Lab Answers
 - Personalized Recommendations
 - Periodic Trends And The Properties Of Elements Lab Answers User Reviews and Ratings

- Periodic Trends And The Properties Of Elements Lab Answers and Bestseller Lists
- 5. Accessing Periodic Trends And The Properties Of Elements Lab Answers Free and Paid eBooks
 - Periodic Trends And The Properties Of Elements Lab Answers Public Domain eBooks
 - Periodic Trends And The Properties Of Elements Lab Answers eBook Subscription Services
 - Periodic Trends And The Properties Of Elements Lab Answers Budget-Friendly Options
- 6. Navigating Periodic Trends And The Properties Of Elements Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends And The Properties Of Elements Lab Answers Compatibility with Devices
 - Periodic Trends And The Properties Of Elements Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends And The Properties Of Elements Lab Answers
 - Highlighting and Note-Taking Periodic Trends And The Properties Of Elements Lab Answers
 - Interactive Elements Periodic Trends And The Properties Of Elements Lab Answers
- 8. Staying Engaged with Periodic Trends And The Properties Of Elements Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends And The Properties Of Elements Lab Answers
- 9. Balancing eBooks and Physical Books Periodic Trends And The Properties Of Elements Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends And The Properties Of Elements Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends And The Properties Of Elements Lab Answers
 - Setting Reading Goals Periodic Trends And The Properties Of Elements Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends And The Properties Of Elements Lab Answers
 - Fact-Checking eBook Content of Periodic Trends And The Properties Of Elements Lab Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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