

7. Chromite is the main ore of chromium. The chromium-containing compound in chromite is FeCr_2O_4 . The percentage of chromium in a sample of chromite can be determined using the method below.

Step 1

A 5.28 g sample of chromite ore is heated with sodium peroxide, Na_2O_2 .



Water is added to the resulting mixture.

Na_2O_2 dissolves in the water forming a solution containing CO_3^{2-} ions.

Step 2

The mixture from Step 1 is filtered and the filtrate is made up to 1.00 dm³ in a volumetric flask.

A 25.0 cm³ sample of this alkaline solution is pipetted into a conical flask and an excess of aqueous potassium iodide is added.

- A redox reaction takes place between I^- ions, CrO_4^{2-} ions and H_2O .
- In this reaction 1 mol CrO_4^{2-} reacts 1.5 mol I_2 .

Step 3

The resulting mixture is titrated with 0.100 mol dm⁻³ sodium thiosulfate, $\text{Na}_2\text{S}_2\text{O}_3$, to estimate the I_2 present.



The average titre of $\text{Na}_2\text{S}_2\text{O}_3$ was 26.5 cm³.

- (a) In Step 1, Na_2O_2 and NaFeO_2 react with water forming an alkaline solution containing a brown precipitate. This is not a redox reaction.

Write equations for:

- the reaction of Na_2O_2 with water
- the reaction of NaFeO_2 with water.



- (b) Calculate the percentage, by mass, of chromium in the ore.

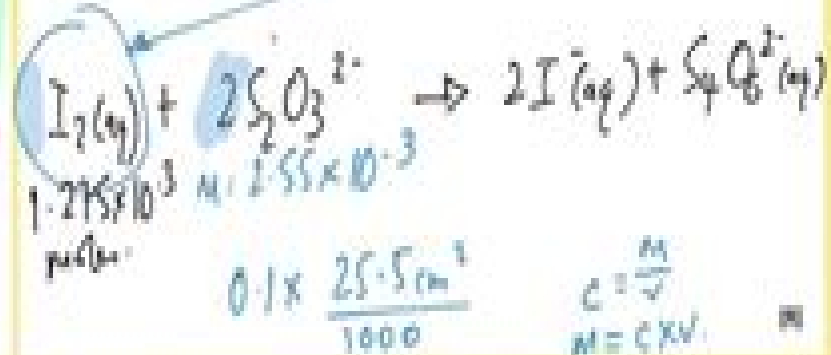
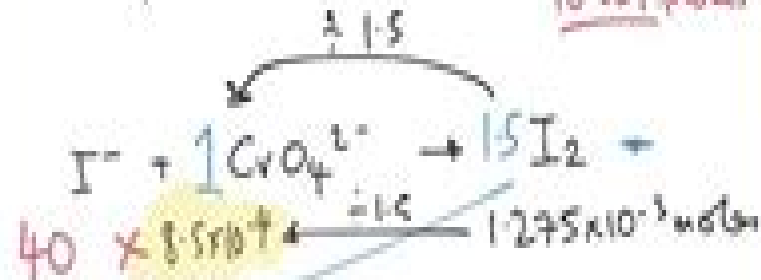
Give your answer to one decimal place.

$$\% \text{ Cr} = \frac{\text{mass of Cr}}{\text{mass of ore}} \times 100$$

$$\% \text{ Cr} = \frac{1.76}{5.28} \times 100 = 33.3\%$$



10.034 moles



- (c) The part refers to Step 2 of the method.

is the redox reaction between I^- ions, CrO_4^{2-} ions and H_2O .

- CrO_4^{2-} ions are reduced to chromate(III) ions.
- I^- ions are oxidised to form I_2 .
- Construct an overall equation for the redox reaction and write half equations for the oxidation and reduction.

Overall equation:

Half equations:

Ocr Chemistry F325 Stretch And Challenge

Mike Smith, John Older



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1. Understanding the eBook Ocr Chemistry F325 Stretch And Challenge
 - The Rise of Digital Reading Ocr Chemistry F325 Stretch And Challenge
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry F325 Stretch And Challenge
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry F325 Stretch And Challenge
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry F325 Stretch And Challenge
 - Personalized Recommendations
 - Ocr Chemistry F325 Stretch And Challenge User Reviews and Ratings
 - Ocr Chemistry F325 Stretch And Challenge and Bestseller Lists

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 - Ocr Chemistry F325 Stretch And Challenge Budget-Friendly Options
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 - Ocr Chemistry F325 Stretch And Challenge Enhanced eBook Features
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 - Highlighting and Note-Taking Ocr Chemistry F325 Stretch And Challenge
 - Interactive Elements Ocr Chemistry F325 Stretch And Challenge
8. Staying Engaged with Ocr Chemistry F325 Stretch And Challenge
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry F325 Stretch And Challenge
9. Balancing eBooks and Physical Books Ocr Chemistry F325 Stretch And Challenge
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry F325 Stretch And Challenge
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry F325 Stretch And Challenge
 - Setting Reading Goals Ocr Chemistry F325 Stretch And Challenge
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry F325 Stretch And Challenge
 - Fact-Checking eBook Content of Ocr Chemistry F325 Stretch And Challenge
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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