

AP Questions: Thermochemistry

1984 B

Substance	Standard Heat of Formation, ΔH_f° in kJ/mol
C(s)	0.00
CO ₂ (g)	-393.5
H ₂ (g)	0.00
H ₂ O(l)	-285.83
O ₂ (g)	0.00
C ₃ H ₇ COOH(l)	?

The enthalpy change for the combustion of butyric acid at 25°C, ΔH_{comb} , is -2,183.5 kilojoules per mole. The combustion reaction is



- From the above data, calculate the standard heat of formation, ΔH_f° , for butyric acid.
- Write a correctly balanced equation for the formation of butyric acid from its elements.

1988 B

Substance	Enthalpy of Combustion, ΔH (kilojoules/mole)
C(s)	-393.5
H ₂ (g)	-285.8
C ₂ H ₅ OH(l)	-1366.7
H ₂ O(l)	--

- Write a separate, balanced chemical equation for the combustion of each of the following: C(s), H₂(g), and C₂H₅OH(l). Consider the only products to be CO₂ and/or H₂O(l).
- In principle, ethanol can be prepared by the following reaction:



Calculate the standard enthalpy change, ΔH , for the preparation of ethanol, as shown in the reaction above.

1988 D

An experiment is to be performed to determine the standard molar enthalpy of neutralization of a strong acid by a strong base. Standard school laboratory equipment and a supply of standardized 1.00 molar HCl and standardized 1.00 molar NaOH are available.

- What equipment would be needed?
- What measurements should be taken?
- Without performing calculations, describe how the resulting data should be used to obtain the standard molar enthalpy of neutralization.
- When a class of students performed this experiment, the average of the results was -55.0 kilojoules per mole. The accepted value for the standard molar enthalpy of neutralization of a strong acid by a strong base is -57.7 kilojoules per mole. Propose two likely sources of experimental error that could account for the result obtained by the class.

1995 B

Propane, C₃H₈, is a hydrocarbon that is commonly used as fuel for cooking.

- Write a balanced equation for the complete combustion of propane gas, which yields CO₂(g) and H₂O(l).
- Calculate the volume of air at 30°C and 1.00 atmosphere that is needed to burn completely 10.0 grams of propane. Assume that air is 21.0 percent O₂ by volume.
- The heat of combustion of propane is -2,220.1 kJ/mol. Calculate the heat of formation, ΔH_f° , of propane given that ΔH_f° of H₂O(l) = -285.3 kJ/mol and ΔH_f° of CO₂(g) = -393.5 kJ/mol.
- Assuming that all of the heat evolved in burning 30.0 grams of propane is transferred to 8.00 kilograms of water (specific heat = 4.18 J/g°C), calculate the increase in temperature of water.

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VM Jensen



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Table of Contents Objectives Questions On Thermochemistry

1. Understanding the eBook Objectives Questions On Thermochemistry
 - The Rise of Digital Reading Objectives Questions On Thermochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Objectives Questions On Thermochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Objectives Questions On Thermochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Objectives Questions On Thermochemistry
 - Personalized Recommendations
 - Objectives Questions On Thermochemistry User Reviews and Ratings
 - Objectives Questions On Thermochemistry and Bestseller Lists
5. Accessing Objectives Questions On Thermochemistry Free and Paid eBooks
 - Objectives Questions On Thermochemistry Public Domain eBooks
 - Objectives Questions On Thermochemistry eBook Subscription Services
 - Objectives Questions On Thermochemistry Budget-Friendly Options
6. Navigating Objectives Questions On Thermochemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Objectives Questions On Thermochemistry Compatibility with Devices
 - Objectives Questions On Thermochemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Objectives Questions On Thermochemistry
 - Highlighting and Note-Taking Objectives Questions On Thermochemistry
 - Interactive Elements Objectives Questions On Thermochemistry
8. Staying Engaged with Objectives Questions On Thermochemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Objectives Questions On Thermochemistry
- 9. Balancing eBooks and Physical Books Objectives Questions On Thermochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Objectives Questions On Thermochemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Objectives Questions On Thermochemistry
 - Setting Reading Goals Objectives Questions On Thermochemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Objectives Questions On Thermochemistry
 - Fact-Checking eBook Content of Objectives Questions On Thermochemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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