## SALT ANALYSIS

## 디

#### SCHEME OF ANALYSIS

EXPERIMENT	OBSERVATION	INFERENCE
PRELIMINARY EXAMINATIONS :		4
1) Noted the colour	Colourless	Absence of Cu2+,Fe2+,Fe3+,Mn2+,Co2+,Ni2+
	Blue	Pre. of Cu <sup>2</sup> *
	Brown	Pre. of Fe <sup>3+</sup>
	Green	Pre. of Ni <sup>2</sup> *
	Pale pink	Pre. of Mn2*
	Pink	Pre. of Co2*
Noted the solubility    a) In water    b) In dil.HCl	a) Soluble and no ppt with Na <sub>2</sub> CO <sub>3</sub> b) Insoluble	a) Pre. of NH <sub>4</sub> * b) Pre. of Pb <sup>2</sup> *
<ol> <li>Flame Test:         <ul> <li>A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.</li> </ul> </li> </ol>	a) Blue surrounded by green b) Brick red c) Crimson red d) Bluish white e) Grassy green on prolonged heating	a) Pre. of Cu <sup>2+</sup> b) Pre. of Ca <sup>2+</sup> c) Pre. of Sr <sup>2+</sup> d) Pre. of Pb <sup>2+</sup> e) Pre. of Ba <sup>2+</sup>
<ol> <li>Ash Test:         <ul> <li>A little of the salt is heated with con.HNO<sub>3</sub> and cobalt nitrate solution. A filter paper is dipped in it and burnt.</li> </ul> </li> </ol>	a) Green ash b) Blue tinted ash c) Pale pink ash	a) Pre. of Zn <sup>2+</sup> b) Pre. of Al <sup>2+</sup> c) Pre. of Mg <sup>2+</sup>
SYSTEMATIC ANALYSIS OF ANIONS	:	
<ol> <li>Dil.H<sub>2</sub>SO<sub>4</sub> test:         To the salt add dil.H<sub>2</sub>SO<sub>4</sub>, warm if required.     </li> </ol>	a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky.	a) Pre. of carbonates
	<ul> <li>b) Smell of vinegar</li> </ul>	b) Pre. of acetate

## **Salt Analysis Manual For P G Course**

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