

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

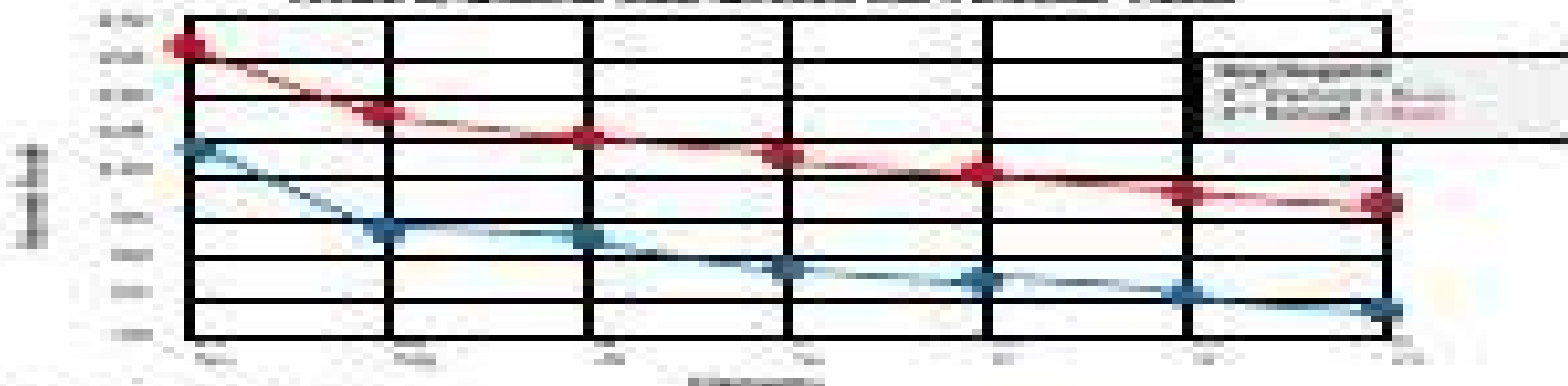
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of a data table/graph below
2. Is a slight downward trend in this graph Period? If not, fill in the value graph
3. Label the line as acid reactivity metallicness
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	1H	2He						
2	3Li	4Be	5B	6C	7N	8O	9F	10Ne
3	11Na	12Mg	13Al	14Si	15P	16S	17Cl	18Ar
4	19K	20Ca	21Sc	22Ti	23V	24Cr	25Mn	26Fe
5	27Co	28Ni	29Cu	30Zn	31Ga	32Ge	33As	34Se
6	37Rb	38Sr	39Y	40Zr	41Nb	42Mo	43Tc	44Ru
7	55Cs	56Ba	57La	58Ce	59Pr	60Nd	61Pm	62Sm

Now label the elements above in the following way:

Trend in Atomic Size Across the Periodic Table



Illustrated Trend Questions

1. Following the trend in atomic size across the periodic table decreases
2. This causes two more than just more reactive properties/behaviours they give off electron giving themselves a negative charge

Worked Trend Questions

1. How would you describe the change in atomic size, i.e. the size of the atom as the periodic movement on your graph shows in relation to the periodic table? decreases
2. How would you describe the change in reactive properties/behaviours (negative and electron) making the electron themselves

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PT Brinkman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the PT Brinkman text.

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