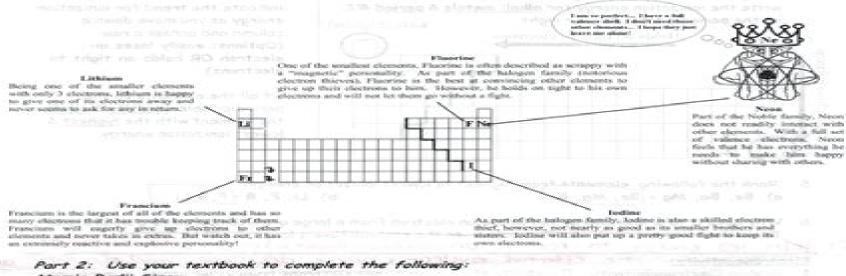
Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

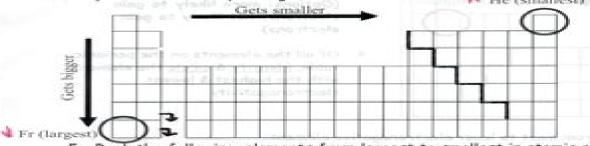
For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)



Atomic Rodii Size:

1. Explain atomic radius: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.

2. Using the information in figure 5-13 (pg 141) in your book, write the atomic radii size for alkali metals & period #2 on the periodic table to your right.



- 3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
- 4. Of all the elements on the periodic table, identify & circle the largest & smallest in size.
- Rank the following elements from largest to smallest in atomic radii: a) Be, Ba, Mg = Ba, Mg, Be
- 6. How can you explain the trend:

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on

in creases the attraction of the nucleus and therefore cloud close to the nucleus resulting

Periodic Trends Pogil Worksheet Answers Extended Answer

Judd E. Hollander

Periodic Trends Pogil Worksheet Answers Extended Answer:

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Table of Contents Periodic Trends Pogil Worksheet Answers Extended Answer

- 1. Understanding the eBook Periodic Trends Pogil Worksheet Answers Extended Answer
 - The Rise of Digital Reading Periodic Trends Pogil Worksheet Answers Extended Answer
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Pogil Worksheet Answers Extended Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Pogil Worksheet Answers Extended Answer
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Pogil Worksheet Answers Extended Answer
 - Personalized Recommendations
 - Periodic Trends Pogil Worksheet Answers Extended Answer User Reviews and Ratings
 - Periodic Trends Pogil Worksheet Answers Extended Answer and Bestseller Lists
- 5. Accessing Periodic Trends Pogil Worksheet Answers Extended Answer Free and Paid eBooks
 - Periodic Trends Pogil Worksheet Answers Extended Answer Public Domain eBooks
 - Periodic Trends Pogil Worksheet Answers Extended Answer eBook Subscription Services
 - Periodic Trends Pogil Worksheet Answers Extended Answer Budget-Friendly Options

- 6. Navigating Periodic Trends Pogil Worksheet Answers Extended Answer eBook Formats
 - o ePub, PDF, MOBI, and More
 - Periodic Trends Pogil Worksheet Answers Extended Answer Compatibility with Devices
 - Periodic Trends Pogil Worksheet Answers Extended Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Periodic Trends Pogil Worksheet Answers Extended Answer
 - Highlighting and Note-Taking Periodic Trends Pogil Worksheet Answers Extended Answer
 - o Interactive Elements Periodic Trends Pogil Worksheet Answers Extended Answer
- 8. Staying Engaged with Periodic Trends Pogil Worksheet Answers Extended Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Pogil Worksheet Answers Extended Answer
- 9. Balancing eBooks and Physical Books Periodic Trends Pogil Worksheet Answers Extended Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Pogil Worksheet Answers Extended Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Pogil Worksheet Answers Extended Answer
 - Setting Reading Goals Periodic Trends Pogil Worksheet Answers Extended Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Pogil Worksheet Answers Extended Answer
 - Fact-Checking eBook Content of Periodic Trends Pogil Worksheet Answers Extended Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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