

## Trends on the Periodic Table

### Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

**Lithium**  
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



**Fluorine**  
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (stoutest electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.

I am so perfect... I have a full electron shell. I don't need those other elements... I hope they just leave me alone!



**Neon**  
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

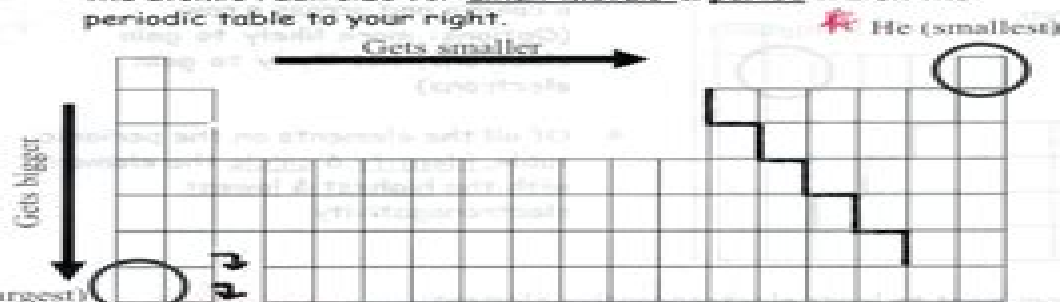
**Francium**  
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

**Iodine**  
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

### Part 2: Use your textbook to complete the following:

#### Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

*Effective nuclear charge,  $Z_{eff}$ , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.*

*not a good enough answer*  
*4AP Answer*

# Periodic Trends Pogil Extension Questions Answers

**Terry C. Jones**



## **Periodic Trends Pogil Extension Questions Answers:**

This book delves into Periodic Trends Pogil Extension Questions Answers. Periodic Trends Pogil Extension Questions Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Periodic Trends Pogil Extension Questions Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Periodic Trends Pogil Extension Questions Answers
- Chapter 2: Essential Elements of Periodic Trends Pogil Extension Questions Answers
- Chapter 3: Periodic Trends Pogil Extension Questions Answers in Everyday Life
- Chapter 4: Periodic Trends Pogil Extension Questions Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Periodic Trends Pogil Extension Questions Answers. The first chapter will explore what Periodic Trends Pogil Extension Questions Answers is, why Periodic Trends Pogil Extension Questions Answers is vital, and how to effectively learn about Periodic Trends Pogil Extension Questions Answers.

3. In chapter 2, the author will delve into the foundational concepts of Periodic Trends Pogil Extension Questions Answers. The second chapter will elucidate the essential principles that need to be understood to grasp Periodic Trends Pogil Extension Questions Answers in its entirety.

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5. In chapter 4, this book will scrutinize the relevance of Periodic Trends Pogil Extension Questions Answers in specific contexts. The fourth chapter will explore how Periodic Trends Pogil Extension Questions Answers is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Periodic Trends Pogil Extension Questions Answers. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Periodic Trends Pogil Extension Questions Answers.

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