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# Ocr B741 02 May 2013 Markscheme

**Zhe Xu**



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web the molar volume will be the volume occupied by one molar mass  $32 \text{ g}$  of the liquid expressing the density in liters instead of  $\text{ml}$  we have  $v = \frac{m}{\rho}$   $32 \text{ g mol}^{-1}$   $790 \text{ g l}^{-1}$   $0.0405 \text{ l mol}^{-1}$  nonnumber

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web the molar volume of a gas from avogadro s law the volume is directly proportional to the number of moles the important point to keep in mind is that  $n$  can be moles of any gas since one mole always contain  $6.02 \times 10^{23}$  molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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volume of gas moles  $V_m$  rearranging this moles of gas  $\frac{V}{V_m}$

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assuming ideal gas behavior at STP the volume of a gas is only dependent on number of moles of

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