How to Identify the Limiting & Excess Reactants/Reagents

- Molar Mass: $C_2H_6 = 30.08 \text{ g/mol}$ $O_2 = 32.00 \text{ g/mol}$ H₂O = 18.02 g/mol
- Calculate how much of one reactant will react with the other reactant
- Ex. What is the limiting reactant when 15.0 g of C2H6 react with 50.0 g of O2?

$$2C_2H_6 + 7O_2 \rightarrow 6H_2O + 4CO_2$$

 15.03 50.03

If reacting all of Reactant A requires more of Reactant B than is available, then B is the Limiting Reactant. If reacting all of Reactant A requires less of Reactant B than is available, then A is the Limiting Reactant.

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