

Oxidation and Reduction

When copper(II) nitrate decomposes, it is transformed to a chemical reaction:

Model 1

Below are chemical equations for reactions that involve copper(II) nitrate as the reactant. Copper gets oxidized in chemical equations involving oxygen gas. Metals react or combine in the presence of an acid solution. Nitric acid (HNO₃) is used in reactions for decomposition and also as an oxidizing agent. In reactions with acids, HNO₃ is often reduced to NO or NO₂. In reactions with acids, HNO₃ is often reduced to NO or NO₂. In reactions with acids, HNO₃ is often reduced to NO or NO₂. In reactions with acids, HNO₃ is often reduced to NO or NO₂.

Model 1 – Redox Reactions

Redox Reactions



Redox Reactions



1. What are the redox reactions in Model 1?

Redox reactions and reduction reactions.

2. Explain the redox and reduction reactions in Model 1. What are the reactants in the redox reactions? What are the products? What are the reactants? What are the products? What are the reactants? What are the products? What are the reactants? What are the products?

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