

PERIODIC CLASSIFICATION OF ELEMENT

DOBERNIER TRIADS:-

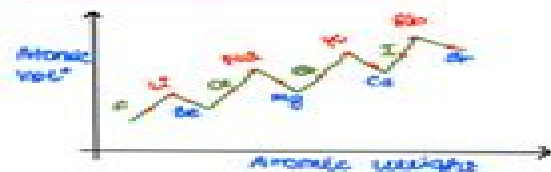
- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



Drawbacks:

- as known elements were not grouped

Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

MODERN PERIODIC TABLE:-

- long form of periodic table
- based on Moseley law: $\sqrt{\lambda} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

Periodicity - Repetition of similar properties of elements after regular interval of time.

Cause of periodicity - due to similar e^- configuration.

Transition elements: they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

Representative elements - Combined elements of s and p block are known as representative elements.

NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga
Eka silicon → Ge

- Corrected atomic mass of element

Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block = G_1 and $G_2 = ns^1-2$

p-Block = $G_{13} - G_{18} = ns^2 np^1-6$

d-Block = $G_3 - G_{10} = (n-1)d^1-10 ns^2$

f-Block = lanthanoid = $2 = 58 - 71$

G_3 and P_6

actinoids = $2 = 90 - 103$

G_3 and P_6

Diagonal Relation:

Li → Be → B → similar
Na → Mg → Al → change ratio
size

Nomenclature:

0 → nil	5 → pent	Suffix = turn
1 → un	6 → hex	100 unnilpentium
2 → bi	7 → sept	
3 → tri	8 → oct	
4 → quad	9 → non	

EFFECTIVE NUCLEAR CHARGE

$Z_{eff} = Z - \sigma$

Trends - Period → $Z_{eff} \uparrow$

Group → Z_{eff} almost remains same. or $Z_{eff} \uparrow \downarrow$

Shielding → $s > p > d > f$

→ more +ve charge, more is Z_{eff}

→ more -ve charge, less is Z_{eff} .

IONISATION ENERGY

- Energy required to remove an electron from an isolated gaseous atom.



IE = IE₁ < IE₂ < IE₃ ...

Trends - $Z_{eff} \propto$ IE

Period - $Z_{eff} \uparrow =$ IE \uparrow

Group - $Z_{eff} \downarrow =$ IE \downarrow

$G_2 = O > AL > Ga > In$

$G_{14} = C > Si > Ge > Pb > Sn$

Period 2:-

$Li < Be < B < C < N < O < F < Ne$

ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

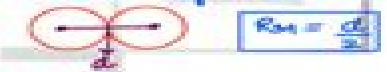
Reason-1 = atoms too small

2 = atom doesn't have any boundary

Covalent Radius - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



Metallic Radius - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



Van der Waals Radius - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



Trends: Atomic Radius $\propto \frac{1}{Z_{eff}}$

Period → $Z_{eff} \uparrow =$ Atomic Radius \downarrow

Group → $Z_{eff} \downarrow =$ Atomic Radius \uparrow

Ionic Radius + Radius < Neutral Radius

ELECTRON GAIN ENTHALPY:-

- Energy required or released when an e^- is added to outermost orbit of an isolated gaseous atom. Is known as Electron gain enthalpy.



Note - In majority cases $\Delta_{eg}H_2$ is -ve (exothermic)

Exception = $\begin{matrix} B \\ N \end{matrix}$ } $\Delta_{eg}H_2 = +ve$ due to stable configuration noble gas

Note - Successive electron gain enthalpy is always +ve.

Trends: $\Delta_{eg}H \propto Z_{eff}$

Period - $Z_{eff} \uparrow \Rightarrow \Delta_{eg}H \uparrow$

Group - $Z_{eff} \downarrow \Rightarrow \Delta_{eg}H \downarrow$

$G_2 = S > Se > Te > Po > O$
→ due to small size Reason!
→ e^-e^- repulsion.

Periodic Classification 3rd Chapter 11th Class

RJ Shavelson



Periodic Classification 3rd Chapter 11th Class:

Periodic Classification 3rd Chapter 11th Class Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Periodic Classification 3rd Chapter 11th Class**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence.

Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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Table of Contents Periodic Classification 3rd Chapter 11th Class

1. Understanding the eBook Periodic Classification 3rd Chapter 11th Class
 - The Rise of Digital Reading Periodic Classification 3rd Chapter 11th Class
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Classification 3rd Chapter 11th Class
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Classification 3rd Chapter 11th Class
 - Personalized Recommendations
 - Periodic Classification 3rd Chapter 11th Class User Reviews and Ratings
 - Periodic Classification 3rd Chapter 11th Class and Bestseller Lists

5. Accessing Periodic Classification 3rd Chapter 11th Class Free and Paid eBooks
 - Periodic Classification 3rd Chapter 11th Class Public Domain eBooks
 - Periodic Classification 3rd Chapter 11th Class eBook Subscription Services
 - Periodic Classification 3rd Chapter 11th Class Budget-Friendly Options
6. Navigating Periodic Classification 3rd Chapter 11th Class eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Classification 3rd Chapter 11th Class Compatibility with Devices
 - Periodic Classification 3rd Chapter 11th Class Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Classification 3rd Chapter 11th Class
 - Highlighting and Note-Taking Periodic Classification 3rd Chapter 11th Class
 - Interactive Elements Periodic Classification 3rd Chapter 11th Class
8. Staying Engaged with Periodic Classification 3rd Chapter 11th Class
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Classification 3rd Chapter 11th Class
9. Balancing eBooks and Physical Books Periodic Classification 3rd Chapter 11th Class
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Classification 3rd Chapter 11th Class
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Classification 3rd Chapter 11th Class
 - Setting Reading Goals Periodic Classification 3rd Chapter 11th Class
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Classification 3rd Chapter 11th Class
 - Fact-Checking eBook Content of Periodic Classification 3rd Chapter 11th Class
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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