

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.8 grams. The actual yield was 23.8 grams. Calculate the percent yield of sodium chloride.

- 2 8.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.8 g. Calculate the percent yield of carbon dioxide.

- 3 1.8 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 11.8 grams. Calculate the percent yield of zinc sulfide.

- 4 1.0 grams of phosphorus reacts with excess chlorine.



The theoretical yield of phosphorus trichloride is 4.0 grams. Calculate the percent yield of phosphorus trichloride.

- 5 Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 6 The actual yield was 18.8 grams. Calculate the percent yield.

- 7 Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2). 1.0 gram of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 8 The actual yield was 40.2 grams. Calculate the percent yield.

- 9 Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 15.8 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} \\ = 0.158 \text{ mol CaCO}_3 \times 1 \text{ mol CaCl}_2 \\ = 0.158 \text{ mol CaCl}_2 \end{array}$$

$$\begin{array}{r} 0.158 \text{ mol CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCl}_2} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} \\ = 0.158 \text{ mol CaCl}_2 \times 146.97 \text{ g CaCl}_2 \\ = 23.22 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \times 2 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

- 10 The actual yield was 22.8 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{22.8 \text{ g}}{23.22 \text{ g}} \times 100 = 9.81\% \times 100 = 981.0\% \text{ CaCl}_2$$

- 11 Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 1.8 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 1.8 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} \\ = 0.046 \text{ mol K} \times 2 \text{ mol KI} = 0.092 \text{ mol KI} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

ANSWER KEY INCLUDED

%
YIELD

Percent Yield Example Answer Key

**Daniel W Foster Professor of Medical
Ethics John Sadler**



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