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G Thomas



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2 p 1 b 1 s 2 2 s 1 2 p 1 1 s 2 2 p 2 c 1 s 2 2 p 2 1 s 2 1 p 2 d 1 s 2 1 p 2

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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potassium element has 19 electrons its electron configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ or simply put at $4s^1$ what is the total number of electrons that can be filled in s p and d subshells

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