

Polyatomic Ions WS

Ammonium (NH_4)¹⁺
 Carbonate (CO_3)²⁻
 Chromate (CrO_4)²⁻

Dichromate (Cr_2O_7)²⁻
 Hydroxide (OH)¹⁻
 Nitrate (NO_3)¹⁻

Nitrite (NO_2)¹⁻
 Sulfate (SO_4)²⁻
 Sulfite (SO_3)²⁻

Phosphate (PO_4)³⁻

Write the NAME of each of the following compounds.

1. $(\text{NH}_4)\text{Cl}$ Ammonium Chloride
2. $\text{Be}(\text{SO}_4)$ Beryllium Sulfate
3. $(\text{NH}_4)_3\text{N}$ Ammonium Nitride
4. MgCl_2 Magnesium Chloride
5. $\text{NH}_4(\text{NO}_3)$ Ammonium Nitrate
6. $\text{Sr}_3(\text{PO}_4)_2$ Strontium Phosphate
7. $\text{Zn}(\text{CrO}_4)_2$ Zinc Dichromate
8. $\text{K}_2(\text{Cr}_2\text{O}_7)$ Potassium Dichromate

9. $\text{Ga}(\text{ClO}_3)_3$ Gallium Chlorate
10. $\text{Cu}(\text{OH})$ Copper I Hydroxide
11. $(\text{NH}_4)_3(\text{PO}_4)$ Ammonium Phosphate
12. $\text{Fe}(\text{SO}_4)$ Iron II Sulfate
13. $\text{Mg}(\text{NO}_2)_2$ Magnesium Nitrite
14. $(\text{NH}_4)\text{NO}_2$ Ammonium Nitrite
15. $\text{Na}_2(\text{Cr}_2\text{O}_7)$ Sodium Dichromate
16. $\text{Na}(\text{OH})$ Sodium Hydroxide

Write the CHEMICAL FORMULA for each of the given NAMES.....cross charges

17. calcium carbonate $\text{Ca}^{2+} \text{ } (\text{CO}_3)^{2-}$ = $\text{Ca}_2(\text{CO}_3)_2$ = CaCO_3
18. barium nitrate $\text{Ba}^{+2} \text{ } \text{NO}_3^{-1}$ = $\text{Ba}(\text{NO}_3)$
19. ammonium sulfate $\text{NH}_4^{+1} \text{ } \text{SO}_4^{-2}$ = $(\text{NH}_4)_2 \text{SO}_4$
20. aluminum hydroxide $\text{Al}^{+3} \text{ } \text{OH}^{-1}$ = $\text{Al}(\text{OH})_3$
21. calcium phosphate $\text{Ca}^{+2} \text{ } \text{PO}_4^{-3}$ = $\text{Ca}_3(\text{PO}_4)_2$
22. cesium nitrate $\text{Cs} \text{ } \text{NO}_2^{-1}$ = CsNO_3
23. sodium nitrite $\text{Na}^{+1} \text{ } \text{NO}_3^{-1}$ = NaNO_3
24. calcium sulfate $\text{Ca}^{+2} \text{ } \text{SO}_4^{-2}$ = CaSO_4
25. beryllium sulfate $\text{Be} \text{ } \text{SO}_4^{-2}$ = BeSO_4
26. sodium carbonate $\text{Na}^{+1} \text{ } \text{CO}_3^{-2}$ = Na_2CO_3
27. magnesium phosphate $\text{Mg}^{+2} \text{ } \text{PO}_4^{-3}$ = $\text{Mg}_3(\text{PO}_4)_2$
28. calcium phosphate $\text{Ca}^{+2} \text{ } \text{PO}_4^{-3}$ = $\text{Ca}_3(\text{PO}_4)_2$

Polyatomic Ions Ws

R Sanford

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Polyatomic Ions Ws

1. Understanding the eBook Polyatomic Ions Ws
 - The Rise of Digital Reading Polyatomic Ions Ws
 - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ions Ws
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polyatomic Ions Ws
 - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ions Ws
 - Personalized Recommendations
 - Polyatomic Ions Ws User Reviews and Ratings
 - Polyatomic Ions Ws and Bestseller Lists
5. Accessing Polyatomic Ions Ws Free and Paid eBooks
 - Polyatomic Ions Ws Public Domain eBooks
 - Polyatomic Ions Ws eBook Subscription Services
 - Polyatomic Ions Ws Budget-Friendly Options

6. Navigating Polyatomic Ions Ws eBook Formats
 - ePub, PDF, MOBI, and More
 - Polyatomic Ions Ws Compatibility with Devices
 - Polyatomic Ions Ws Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Polyatomic Ions Ws
 - Highlighting and Note-Taking Polyatomic Ions Ws
 - Interactive Elements Polyatomic Ions Ws
8. Staying Engaged with Polyatomic Ions Ws
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polyatomic Ions Ws
9. Balancing eBooks and Physical Books Polyatomic Ions Ws
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Polyatomic Ions Ws
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Polyatomic Ions Ws
 - Setting Reading Goals Polyatomic Ions Ws
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Polyatomic Ions Ws
 - Fact-Checking eBook Content of Polyatomic Ions Ws
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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