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# FORMULAS WITH POLYTOMIC IONS

Matching the horizontal and vertical axes, write the formulas of the compounds with the following combination of ions.

	$\text{OH}^-$	$\text{OH}_3^-$	$\text{CO}_3^{+2}$	$\text{SO}_4^{+2}$	$\text{PO}_4^{+2}$
$\text{H}^+$	$\text{HOH}$ $(\text{H}_2\text{O})$	$\text{HNO}_3$	$\text{H}_2\text{CO}_3$	$\text{H}_2\text{SO}_4$	$\text{H}_3\text{PO}_4$
$\text{Na}^+$	$\text{NaOH}$	$\text{NaNO}_3$	$\text{Na}_2\text{CO}_3$	$\text{Na}_2\text{SO}_4$	$\text{Na}_3\text{PO}_4$
$\text{Mg}^{+2}$	$\text{Mg}(\text{OH})_2$	$\text{Mg}(\text{NO}_3)_2$	$\text{MgCO}_3$	$\text{MgSO}_4$	$\text{Mg}_3(\text{PO}_4)_2$
$\text{NH}_4^{+1}$	$\text{NH}_4\text{OH}$	$\text{NH}_4\text{NO}_3$	$(\text{NH}_4)_2\text{CO}_3$	$(\text{NH}_4)_2\text{SO}_4$	$(\text{NH}_4)_3\text{PO}_4$
$\text{Ca}^{+2}$	$\text{KOH}$	$\text{KNO}_3$	$\text{K}_2\text{CO}_3$	$\text{K}_2\text{SO}_4$	$\text{K}_3\text{PO}_4$
$\text{K}^+$	$\text{Al}(\text{OH})_3$	$\text{Al}(\text{NO}_3)_3$	$\text{Al}_2(\text{CO}_3)_3$	$\text{Al}_2(\text{SO}_4)_3$	$\text{AlPO}_4$
$\text{Al}^{+3}$	$\text{Pb}(\text{OH})_4$	$\text{Pb}(\text{NO}_3)_4$	$\text{Pb}(\text{CO}_3)_2$	$\text{Pb}(\text{SO}_4)_2$	$\text{Pb}_3(\text{PO}_4)_4$

# Polyatomic Ionic Formulas Answers

**Stanley E. Manahan**



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