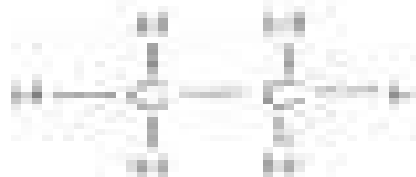


MICHEL 1: Relative electronegativities: $\text{O} > \text{N} > \text{C} > \text{H}$

A



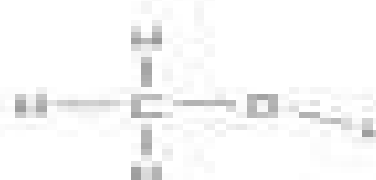
B



C



D



The bonds in B and C are polar.
 The molecules A and D are non-polar.

QUESTIONS

- For each molecule indicated 1, draw in the polar covalent bonds between the atoms. Indicate which electrons are shared and which.

 - It may prove to help distinguishing the nature of the bonds that electrons share in the molecule covalent bonds
 - It both atoms are equally electronegative, draw the electron in the middle of the bond

- For each molecule, draw the distribution of all bonding electrons that are shared covalently. Draw either a partial arrangement to show bonds, or or or
- Label all polar bonds with an arrow (↗).
- Label a polar bond as a dipole bond. Does the first atom have a partial positive charge or a partial negative charge? Give the partial positive/negative bond and give the partial charge (δ+ or δ-).
- Class molecules that are polar and non-polar. Give the polar molecules (indicate δ+ or δ-). Give the non-polar molecules (indicate δ+ or δ-).

Pogil Biological Molecules Answer Key

ML Morrison



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