

INTERMOLECULAR FORCES – ANSWER KEY (from Tro, Chapter 11, page 512)

49. Determine the kinds of intermolecular forces that are present in each of the following elements or compounds:

- | | | | |
|----------------------|--------------------------|------------------------|------------------------|
| (a) Kr* | (b) NCl ₃ ** | (c) SiH ₄ * | (d) HF**** |
| (e) N ₂ * | (f) NH ₃ **** | (g) CO** | (h) CCl ₄ * |

50. Determine the kinds of intermolecular forces that are present in each of the following elements or compounds:

- | | | | |
|------------------------|--------------------------|----------------------------|----------------------|
| (a) HCl** | (b) H ₂ O**** | (c) Br ₂ * | (d) He* |
| (e) PH ₃ ** | (f) HBr** | (g) CH ₃ OH**** | (h) I ₂ * |

* dispersion forces only; ** dispersion and dipole-dipole forces; **** dispersion force, dipole-dipole forces, and hydrogen bonds

51. Arrange the following in order of increasing boiling point. Explain your reasoning.

- | | | | |
|---------------------|-------------------------------------|--|---|
| (a) CH ₄ | (b) CH ₃ CH ₃ | (c) CH ₃ CH ₂ Cl | (d) CH ₃ CH ₂ OH a b c d |
|---------------------|-------------------------------------|--|---|

52. Arrange the following in order of increasing boiling point. Explain your reasoning.

- | | | |
|----------------------|-----------------------|-----------------------------------|
| (a) H ₂ S | (b) H ₂ Se | (c) H ₂ O a b c |
|----------------------|-----------------------|-----------------------------------|

53. For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning.

- | | |
|--|---|
| (a) CH₃OH or CH ₃ SH | (b) CH ₃ OCH ₃ or CH₃CH₂OH |
| (c) CH ₄ or CH₃CH₃ | |

54. For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning.

- | | | |
|--|--|--|
| (a) NH₃ or CH ₄ | (b) CS₂ or CO ₂ | (c) CO ₂ or NO₂ |
|--|--|--|

55. For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning.

- | | | |
|---|---|--|
| (a) Br₂ or I ₂ | (b) H₂S or H ₂ O | (c) NH ₃ or PH₃ |
|---|---|--|

56. For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning.

- | |
|---|
| (a) CH₄ or CH ₃ Cl |
| (b) CH ₃ CH ₂ CH ₂ OH or CH₃OH |
| (c) CH ₃ OH or H₂CO |

57. Which of the following pairs of substances would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved.

- | | |
|---|--|
| (a) CCl ₄ and H ₂ O | (b) KCl and H₂O (ion-dipole forces) |
| (c) Br₂ and CCl₄ (dispersion forces) | (d) CH₃CH₂OH and H₂O (H-bonds) |

58. Which of the following pairs of compounds would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved.

- | |
|--|
| (a) CH₃CH₂CH₂CH₂CH₃ and CH₃CH₂CH₂CH₂CH₂CH₃ (dispersion forces) |
| (b) CBr ₄ and H ₂ O |
| (c) LiNO₃ and H₂O (ion-dipole forces) |
| (d) CH ₃ OH and CH ₃ CH ₂ CH ₂ CH ₂ CH ₃ |

Pogil Answer Key Intermolecular

Marcel A. Müller



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Table of Contents Pogil Answer Key Intermolecular

1. Understanding the eBook Pogil Answer Key Intermolecular
 - The Rise of Digital Reading Pogil Answer Key Intermolecular
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Answer Key Intermolecular
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Answer Key Intermolecular
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Answer Key Intermolecular
 - Personalized Recommendations

- Pogil Answer Key Intermolecular User Reviews and Ratings
- Pogil Answer Key Intermolecular and Bestseller Lists
- 5. Accessing Pogil Answer Key Intermolecular Free and Paid eBooks
 - Pogil Answer Key Intermolecular Public Domain eBooks
 - Pogil Answer Key Intermolecular eBook Subscription Services
 - Pogil Answer Key Intermolecular Budget-Friendly Options
- 6. Navigating Pogil Answer Key Intermolecular eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Answer Key Intermolecular Compatibility with Devices
 - Pogil Answer Key Intermolecular Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Answer Key Intermolecular
 - Highlighting and Note-Taking Pogil Answer Key Intermolecular
 - Interactive Elements Pogil Answer Key Intermolecular
- 8. Staying Engaged with Pogil Answer Key Intermolecular
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Answer Key Intermolecular
- 9. Balancing eBooks and Physical Books Pogil Answer Key Intermolecular
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Answer Key Intermolecular
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Answer Key Intermolecular
 - Setting Reading Goals Pogil Answer Key Intermolecular
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Answer Key Intermolecular
 - Fact-Checking eBook Content of Pogil Answer Key Intermolecular

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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