#### POGIL 2 Ions

How are lone made from neutral atoms?

You have formed that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore atomic variation that can take place—the loss and gain af electrons. The exchange of electrons between atomics a very common way for chemical change to take place. We will see it many times throughout the year.

Use Disserum 1 or the next asset to complete the following: table.

	Metal or Nonmetal	Number of pretens inthe store	Wassber of protons in the ion	Mumber of neutrons in the attent	Number of secotroes in the ion	Number of electrons in the atom	Number of electrons in the item	Change on the lon
Lithium:	rected	3		4	-	3	2	13%
Magnissians	Africal	12	7.2	7.2	12	12	1.0	200
Akaminum	Metalloid	13	7.3	14	34	15	10	311
Physicae	Non-metal	9	and the same	10	10	0	1.0	100
Ovegen	(Spirestone)	8	N		.81	8	1.0	2-
Nitrogen	Non-reseted	7	7	7	7	7	5.0	3-

- Based on the table you completed in Question 1, what distinguishes a neutral atom flore an ion?
   A weatral atom has the same number of protons and electrons (no charge). An ion has a charge (it of protons does not equal it of electrons).
- 3. Examine the isstope symbols in Diagram I...
  - a. Where is the lan charge located in the leasupe symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located? No.
- Which substrettés particle carries a positive charge?

Protous are positive

5. Which subspenie particle carries a negative charge?

Electrons are negative

 Propose a mathematical equation to calculate the charge on an ion from the number of pretons and electrons in an ion. Confirm that your equation works using two posts valors examples and two possitive ion examples from Disserse: 1.

The charge of an ion = 0 of protons - 8 of electrons

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