

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

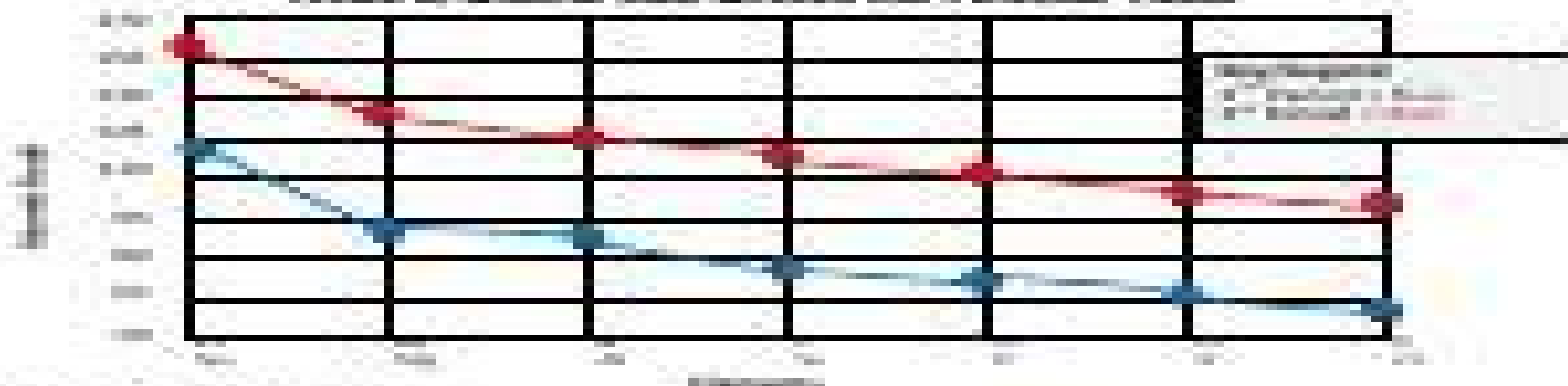
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Draw what graph (ionic radii, electronegativity, etc) data points graphed below
2. Is a slight downward trend in this graph Period? If not, is the data graph
3. Label each box as ionization energy
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	Li	Be	B	C	N	O	F	Ne
2	Na	Mg	Al	Si	P	S	Cl	Ar
3	K	Ca	Sc	Ti	V	Cr	Mn	Fe
4	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru
5	Cs	Ba	La	Hf	Ta	W	Re	Os
6	Fr	Ra	Ac	Th	Pa	U	Np	Pu

Now label the elements above in the following way:

Trend in Atomic Size across the Periodic Table



Electronegativity Trend Equations

1. Following the trend in atomic size across the periodic table, electronegativity
2. This means two atoms that are more electronegative (higher) will have a stronger pulling force on the bonding electrons.

Electronegativity Trend Equations

1. Electronegativity trend in the periodic table is downward (e.g. Na, K, Rb, Cs, Fr has a decreasing electronegativity as you go down in columns of the periodic table)
2. This means two atoms that are more electronegative (higher) will have a stronger pulling force on the bonding electrons.

Periodicity Activity Okstate Answer Key

Brendan G. Carr



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