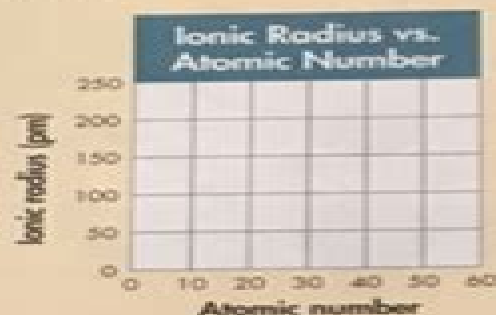


Quick Lab

Purpose To use a graph to identify period and group trends

Materials

- graph paper
- pencil



Periodic Trends in Ionic Radii

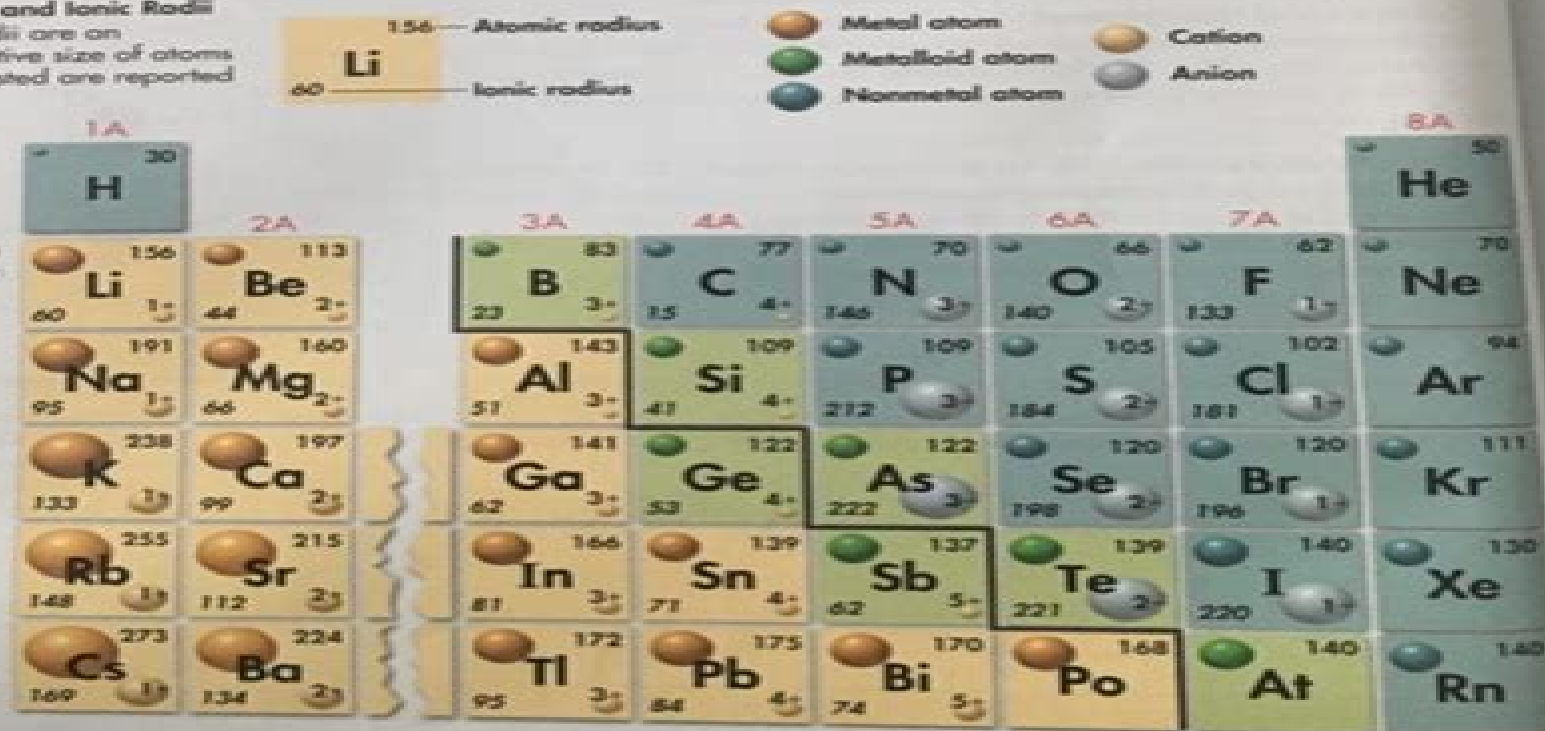
Procedure

Use the data presented in Figure 6.23 to plot ionic radius versus atomic number.

Analyze and Conclude

- 1. Compare** How does the size change when an atom forms a cation and when an atom forms an anion?
- 2. Describe** How do the ionic radii vary within a group of metals? How do they vary within a group of nonmetals?
- 3. Describe** What is the shape of a portion of the graph that corresponds to one period?
- 4. Compare and Contrast** Is the trend across a period similar or different for Periods 2, 3, 4, and 5?
- 5. Explain** Propose explanations for the trends you have described for ionic radii within groups and across periods.

Figure 6.23 Atomic and Ionic Radii
Atomic and ionic radii are an indication of the relative size of atoms and ions. The data listed are reported in picometers (pm).



Periodic Trends In Ionic Radii Lab Answers

**Charlie Harding, David Arthur
Johnson, Rob Janes**



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Table of Contents Periodic Trends In Ionic Radii Lab Answers

1. Understanding the eBook Periodic Trends In Ionic Radii Lab Answers
 - The Rise of Digital Reading Periodic Trends In Ionic Radii Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends In Ionic Radii Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends In Ionic Radii Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends In Ionic Radii Lab Answers
 - Personalized Recommendations
 - Periodic Trends In Ionic Radii Lab Answers User Reviews and Ratings
 - Periodic Trends In Ionic Radii Lab Answers and Bestseller Lists

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 - Periodic Trends In Ionic Radii Lab Answers Public Domain eBooks
 - Periodic Trends In Ionic Radii Lab Answers eBook Subscription Services
 - Periodic Trends In Ionic Radii Lab Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Periodic Trends In Ionic Radii Lab Answers Compatibility with Devices
 - Periodic Trends In Ionic Radii Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends In Ionic Radii Lab Answers
 - Highlighting and Note-Taking Periodic Trends In Ionic Radii Lab Answers
 - Interactive Elements Periodic Trends In Ionic Radii Lab Answers
8. Staying Engaged with Periodic Trends In Ionic Radii Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends In Ionic Radii Lab Answers
9. Balancing eBooks and Physical Books Periodic Trends In Ionic Radii Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends In Ionic Radii Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends In Ionic Radii Lab Answers
 - Setting Reading Goals Periodic Trends In Ionic Radii Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends In Ionic Radii Lab Answers
 - Fact-Checking eBook Content of Periodic Trends In Ionic Radii Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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