

## 5.1 Atomic structure and the Periodic Table

You will learn:

- To understand how the position of an element in the Periodic Table can be used to predict its atomic structure and chemical and physical properties.
- To understand that models and analogies can change according to scientific evidence
- To describe the uses, strengths and limitations of models and analogies
- To represent scientific ideas using recognised symbols or formulae

1.

Draw a line to match the word with its correct meaning.

The first line has been drawn for you.

| Word      | Meaning   |
|-----------|---|
| Electrons | Particles with no charge found in an atom                 |
| Neutrons  | The central part of the atom                              |
| Protons   | The paths or orbits that electrons move along in the atom |
| Nucleus   | Very small negatively charged particles in an atom        |
| Shells    | Particles with a positive charge found in an atom         |

2.

Complete the sentences.

Use the words in the box.

|        |        |          |           |         |
|--------|--------|----------|-----------|---------|
| atomic | mass   | protons  | electrons | nucleus |
|        | shells | neutrons | atom      |         |

- a An \_\_\_\_\_ number is the protons an \_\_\_\_\_ of an element has.
- b A \_\_\_\_\_ number is the total number of \_\_\_\_\_ and \_\_\_\_\_ found in the nucleus of an atom.

# Periodic Table Workbook Answers Chemistry

**J Rink**



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