

## Part 2

# GCSE Chemistry OCR A (Combined Science - **HIGHER**) Paper 2 Exam Walkthrough

### Combined Science Higher - Chemistry

### Walkthrough Mark Paper

12  
(5)

Look at the equation for the equilibrium reaction between sulfur dioxide,  $\text{SO}_2$ , oxygen and sulfur trioxide,  $\text{SO}_3$ , in a closed system.



- (a) Predict the effect of adding more oxygen to the equilibrium mixture.

Explain your answer.

(2)

- (b) Predict the effect of increasing the total pressure in the equilibrium mixture.

Explain your answer.

(2)

- (c) Predict the effect of increasing the **temperature** of the equilibrium mixture.

Explain your answer.

(2)

- (d) Sulfur dioxide for this reaction is made by burning sulfur.



Calculate the mass of sulfur needed to make 48 tonnes of sulfur dioxide.

Give your answer to 2 significant figures.

The relative atomic mass,  $A_r$ , of S is 32.1 and of O is 16.0.

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**S Nieto**



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