

- (a) N<sub>2</sub>
- (b) O<sub>2</sub>
- (c) Cl<sub>2</sub>
- (d) S

**10. Which gas is used to manufacture king of chemicals sulphuric acid?**

- (a) N<sub>2</sub>
- (b) O<sub>2</sub>
- (c) Cl<sub>2</sub>
- (d) S

**11. Equilibrium constant has no unit when number of moles of reactants and products are;**

- (a) same
- (b) different
- (c) both a & b
- (d) none

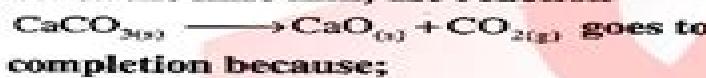
**12. For reactions having large K<sub>c</sub> value, the reaction proceeds to;**

- (a) completion
- (b) equilibrium state
- (c) back ward
- (d) None

**13. The characteristics of reversible reactions are the following except;**

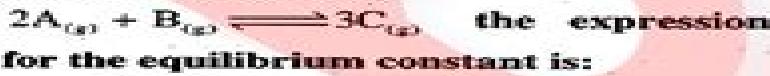
- (a) products never recombine to form reactants
- (b) they never complete
- (c) they proceed in both ways
- (d) they have a double arrow between reactants and products

**14. In the lime kiln, the reaction**



- (a) of high temperature
- (b) CaO is more stable than CaCO<sub>3</sub>
- (c) CO<sub>2</sub> escapes continuously
- (d) CaO is not dissociated

**15. For the reaction,**



- (a)  $\frac{[2\text{A}][\text{B}]}{[3\text{C}]}$
- (b)  $\frac{[\text{A}]^2 [\text{B}]}{[\text{C}]^3}$

$$(c) \frac{[3\text{C}]}{[2\text{A}][\text{B}]} \quad (d) \frac{[\text{C}]^3}{[\text{A}]^2 [\text{B}]}$$

**16. When a system is at equilibrium states?**

- (a) the concentration of reactants and products becomes equal
- (b) the opposing reactions (forward and reverse) stop
- (c) the rate of the reverse reaction becomes very low
- (d) the rates of the forward and reverse reactions becomes equal.

**17. Which one of the following statements is not correct about active mass?**

- (a) rate of reaction is directly proportional to active mass.
- (b) active mass is taken in molar concentration
- (c) active mass is represented by square brackets
- (d) active mass means total mass of substances

**18. When the magnitude of K<sub>c</sub> is very large it indicates;**

- (a) reaction mixture consists of almost all products
- (b) reaction mixture has almost all reactants
- (c) reaction has not gone to completion
- (d) reaction mixture has negligible products

**19. When the magnitude of K<sub>c</sub> is very small it indicates;**

- (a) equilibrium will never establish

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**L Manion**

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