

# PERIODIC CLASSIFICATION OF ELEMENT

## DOBERNIER TRIADS:-

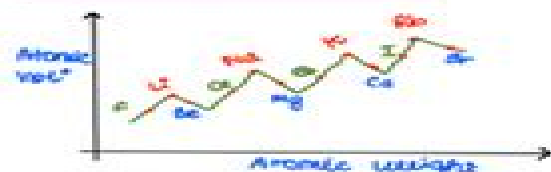
- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



### Drawbacks:-

- as known elements were not grouped

## Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

## MODERN PERIODIC TABLE:-

- long form of periodic table
- based on Moseley law:  $\sqrt{\lambda} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

**Periodicity** - Repetition of similar properties of elements after regular interval of time.

**Cause of periodicity** - due to similar  $e^-$  configuration.

**Transition elements**; they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

**Representative elements** - Combined elements of s and p block are known as representative elements.

## NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



### Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

## MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

### Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga  
Eka silicon → Ge

- Corrected atomic mass of element

### Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

## MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block =  $G_1$  and  $G_2 = ns^1-2$

p-Block =  $G_{13} - G_{18} = ns^2 np^1-6$

d-Block =  $G_3 - G_{10} = (n-1)d^1-10 ns^2$

f-Block = lanthanoid =  $2 = 5f^1-7$

actinoids =  $2 = 5f^1-10$

### Diagonal Relation;

Li → Be → B → similar  
Na → Mg → Al → change ratio  
size

### Nomenclature;

|          |          |                     |
|----------|----------|---------------------|
| 0 → nil  | 5 → pent | Suffix = turn       |
| 1 → un   | 6 → hex  | 100<br>unnilpentium |
| 2 → bi   | 7 → sept |                     |
| 3 → tri  | 8 → oct  |                     |
| 4 → quad | 9 → inn  |                     |

## EFFECTIVE NUCLEAR CHARGE

$Z_{eff} = Z - \sigma$

Trends - Period →  $Z_{eff} \uparrow$

Group →  $Z_{eff}$  almost remains same. or  $Z_{eff} \uparrow \downarrow$

Shielding →  $s > p > d > f$

→ more +ve charge, more is  $Z_{eff}$

→ more -ve charge, less is  $Z_{eff}$ .

## IONISATION ENERGY

→ Energy required to remove an electron from an isolated gaseous atom.

Mg →  $M^{+}_{(g)} + e^-$  IE<sub>1</sub>

$M^{+}_{(g)} \rightarrow M^{2+}_{(g)} + e^-$  IE<sub>2</sub>

IE = IE<sub>1</sub> < IE<sub>2</sub> < IE<sub>3</sub> ...

Trends -  $Z_{eff} \propto$  IE

Period -  $Z_{eff} \uparrow =$  IE  $\uparrow$

Group -  $Z_{eff} \downarrow =$  IE  $\downarrow$

$G_{13} = G > Al > Ga > In$   
 $G_{14} = C > Si > Ge > Pb > Sn$

## Period 2:-

$Li < Be < B < C < N < O < F < Ne$

## ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

Reason-1 = atoms too small

2 = atom doesn't have any boundary

**Covalent Radius** - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



**Metallic Radius** - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



**Van der Waals Radius** - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



Trends; Atomic Radius  $\propto \frac{1}{Z_{eff}}$   
Period →  $Z_{eff} \uparrow =$  Atomic Radius  $\downarrow$   
Group →  $Z_{eff} \downarrow =$  Atomic Radius  $\uparrow$

Ionic Radius + Radius < Neutral Radius

## ELECTRON GAIN ENTHALPY:-

→ Energy required or released when an  $e^-$  is added to outermost orbit of an isolated gaseous atom. Is known as Electron gain enthalpy.



Note - In majority cases  $\Delta_{eg}H_2$  is -ve (exothermic)

Exception = N }  $\Delta_{eg}H_2 = +ve$   
 noble gas } due to stable configuration

Note - Successive electron gain enthalpy is always +ve.

Trends;  $\Delta_{eg}H \propto Z_{eff}$

Period -  $Z_{eff} \uparrow \Rightarrow \Delta_{eg}H \uparrow$

Group -  $Z_{eff} \downarrow \Rightarrow \Delta_{eg}H \downarrow$

$G_{15} = S > Se > Te > Po > O$   
→ due to small size Reason!  
→  $e^-e^-$  repulsion.

# Periodic Classification 3rd Chapter 11th Class

**AN Whitehead**



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