



## Percent Yield

45 grams of nitrogen monoxide (NO) react with an excess of sodium hydroxide (NaOH) following the balanced equation above. 6 grams of water were obtained from the reaction

What is the % yield?

Nitrogen monoxide (NO) gmm:

$$(1)(14.0) + (1)(16.0) = 30.0 \text{ g/mol}$$

$$\text{Actual Yield} = 6 \text{ g H}_2\text{O}$$

Water (H<sub>2</sub>O) gmm:

$$(2)(1.0) + (1)(16.0) = 18.0 \text{ g/mol}$$

$$\text{Theoretical Yield} = 6.75 \text{ g H}_2\text{O}$$

$$\frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100 = 88.9\%$$

$$\frac{45 \text{ g NO} \times \frac{1 \text{ mol NO}}{30 \text{ g NO}} \times \frac{4 \text{ mol NO}}{4 \text{ mol NO}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \times 18 \text{ g H}_2\text{O}}{1} = 6.75 \text{ g H}_2\text{O}$$

## Percent Yield Problems Answers

**EW Minium**



## Percent Yield Problems Answers:

*Dr. Salm's Chemistry Problem Drill Book* Robert J. Salm, Robert J. Salm Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests [CliffsNotes Chemistry Practice Pack](#) Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level [CliffsStudySolver: Chemistry](#) Charles Henrickson, 2007-05-03 The CliffsStudySolver workbooks combine 20 percent review material with 80 percent practice problems and the answers to help make your lessons stick CliffsStudySolver Chemistry is for students who want to reinforce their knowledge with a learn by doing approach Inside you'll get the practice you need to learn Chemistry with problem solving tools such as Clear concise reviews of every topic Practice problems in every chapter with explanations and solutions A diagnostic pretest to assess your current skills A full length exam that adapts to your skill level A glossary examples of calculations and equations and situational tasks can help you practice and understand chemistry This workbook also covers measurement chemical reactions and equations and matter elements compounds and mixtures Explore other aspects of the language including Formulas and ionic compounds Gases and the gas laws Atoms The mole elements and compounds Solutions and solution concentrations Chemical bonding Acids bases and buffers Practice makes perfect and whether you're taking lessons or teaching yourself CliffsStudySolver guides can help you make the grade *Concepts And Problems In Physical Chemistry* P.S. Raghavan, 1997 Contents Introduction Atoms Molecules and Formulas Chemical Equations and Stoichiometry Aqueous Reactions and Solution Stoichiometry Gases Intermolecular Forces Liquids and Solids Atoms Structure and the Periodic Table Chemical Bonding Chemical Thermodynamics Solutions Chemical Kinetics Chemical Equilibrium Acids and Bases Ionic Equilibria I Ionic Equilibria II Redox Reactions Electrochemistry Nuclear Chemistry **Chemistry: 1001 Practice Problems For Dummies (+ Free Online Practice)** Heather Hattori, Richard H. Langley, 2022-05-10 Practice your way to a better grade in your Chemistry class Chemistry 1001 Practice Problems For Dummies gives you 1 001 opportunities to practice solving problems on all the topics covered in your chemistry class in the book and online Get extra practice with tricky subjects

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web 1 set the level at a convenient point 2 perform the temporary adjustments 3 hold the staff vertically over the benchmark observe the staff reading and write in the b s column of the field book 4 hold the staff at other points and note the staff reading in the i s column of the field book 5

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web seventh power in arithmetic and algebra the seventh power of a number  $n$  is the result of multiplying seven instances of  $n$  together so  $n^7 = n \cdot n \cdot n \cdot n \cdot n \cdot n \cdot n$  seventh powers are also formed by multiplying a number by its sixth power the square of a number by its fifth power or the cube of a number by its fourth power the

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web  $7 \times 7 \times 7 \times 7$  can be written as  $7^4$  this is read as seven to the power of four 2 of 4 power notation is used when finding the area of a square or the volume of a cube 16 is a square number

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