

QUESTION 3

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a 0, 2 mol.dm⁻³ ethanoic acid solution to 10 g of NaHCO₃ to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of CO₂ produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm³ of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3 25 cm³ of a potassium hydroxide solution of concentration 0,25 mol.dm⁻³ completely neutralises a dilute solution of sulphuric acid (H₂SO₄) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

[25]

Physical Science Grade 11 Controlled Test 2 September 2014

R Sandford



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Sports Engineering and Computer Science Qi Luo, 2015-05-18 Sports Engineering and Computer Science contains papers presented at the 2014 International Conference on Sport Science and Computer Science SSCS 2014 held September 16 17 2014 in Singapore and at the 2014 International Conference on Biomechanics and Sports Engineering BSE 2014 held October 24 25 2014 in Riga Latvia The contributions have

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