

Photoelectron Spectroscopy Worksheet

<http://www.scribd.com/doc/91590491/9/Photoelectron-Spectroscopy>

1. In a photoelectron spectrum, photons of 165.7 MJ/mol strike atoms of an unknown element. If the kinetic energy of the ejected electrons is 25.4 MJ/mol, what is the ionization energy of the element?

$$165.7 \text{ MJ/mol} - 25.4 \text{ MJ/mol} = 140.3 \text{ MJ/mol}$$

2. What determines the position and the height (intensity) of each peak in a photoelectron spectrum?

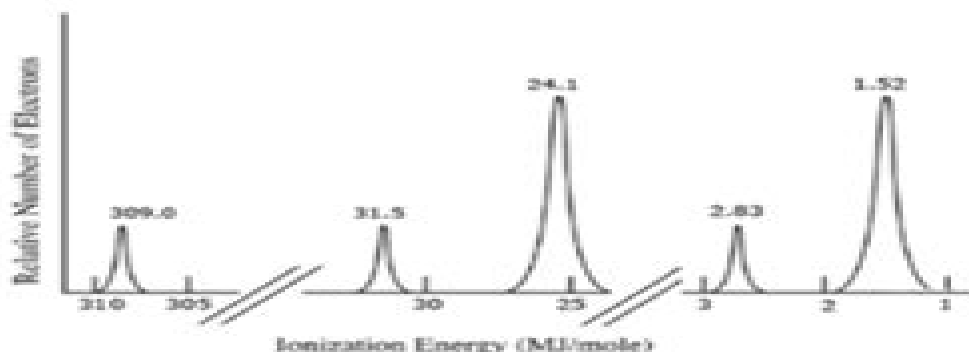
The position of each peak is identified by the ionization energy, the height of each peak identifies the ratio of electrons in each level or orbital.

3. Why is the distance of the energy level from the nucleus important in determining the corresponding peak position in the photoelectron spectrum?

The closer the electron, the more interaction and attraction, the more energy is needed to get rid of that electron.

4. Based on the information provided below, draw a photoelectron spectrum for argon. Indicate the relative intensities and positions of all peaks.

$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$
- 309.0	- 31.5	- 24.1	- 2.83	- 1.52



Photoelectron Spectrum Ap Chemistry Questions

H Kauffman

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