

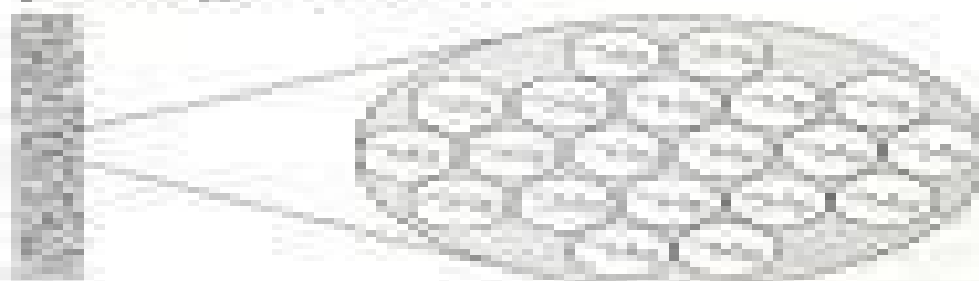
Average Atomic Mass

How can the masses of the particles tell us something?

Why?

Many different isotopes exist. They are naturally occurring species. As you learned previously, the masses of these isotopes have different values. However, regardless of particular isotopic abundances in the natural element, the mass from different isotopes contributes equal amounts of particles and therefore equal mass. Hence, you cannot simply take the mass of a particular isotope and calculate what the mass of the element would be. You must take into account the relative amount of each isotope that contributes to the element's mass.

Model 1 – A Strip of Magnesium Metal



1. What are the particles making up metal Mg shown in Model 1?
They are atoms shown below in Model 2.
2. What are the mass numbers of the isotopes comprising samples of magnesium shown in Model 1?
24, 25, 26
3. How all of the masses of magnesium in Model 1 have the same positive number? Explain.
All have the same number of protons, but different electron masses and shells.
4. How the sample of 24 isotopes of magnesium shown in Model 1 is shown to represent the mass numbers of all of the isotopes and the relative amount of each isotope present.

	^{24}Mg	^{25}Mg	^{26}Mg
Number of atoms	10	5	5

5. Calculate masses of magnesium in the mass numbers in Model 1.
 ^{24}Mg
6. Model in Model 1 shows the sample composed of 24 atoms of ^{24}Mg , 5 atoms of ^{25}Mg , and 5 atoms of ^{26}Mg . How does this model represent the average mass of the element?
As a group of 24 magnesium atoms, there will be approximately one ^{24}Mg , one ^{25}Mg , and eight ^{26}Mg atoms.

Periodicity Of Elements Pogil Key

Tom Jackson



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