

Periodicity Lab

HS-PS1-1

PERIODICITY LAB

Lesson Duration: 60 Minutes + 10 Introductory Lab Prep

Description

This lab has students looking at the reactivity of alkaline earth metals with sulfate using well plates. Students will be creating various solutions of alkaline earth metal solutions and testing these solutions to gauge relative reactivity at different concentrations. In addition to doing the groups for each element. After collecting their data, they will answer various lab analysis questions about reactivity trends of metals, draw their models, and determine how this trend would differ for non-metals.

Use the included Periodicity Lab Prep Sheet to put your lab together!

This product comes with an answer key!

NGSS Alignment:

Performance Expectation: HS-PS1-1 Use the periodic table as a model to predict the relative reactivity of elements based on the patterns of electrons in the outermost energy level of atoms.

Disciplinary Core Idea: PS1-4 Structure and Properties of Matter: Every atom has a charged substructure consisting of a nucleus, which is made of protons and neutrons, surrounded by electrons. The periodic table orders elements horizontally by the number of protons in the atoms' nucleus and groups those with similar chemical properties in columns. The repeating patterns of this table reflect patterns of outer electron shells.

Cross-Cutting Concepts: Different patterns may be observed at each of the scales at which a system is studied and can provide evidence for causality in explanations of phenomena.

Science and Engineering Practices: Use a model to predict the relationships between systems or between components of a system.

Lesson Objectives

Students will be able to...

1. Identify trends in reactivity of metals and non-metals in the periodic table.

Name _____ Date _____ Period _____

Periodicity Lab Why?

The periodic table can be used to predict the physical and chemical properties of elements and their compounds. This lab activity investigates the reactivity of group 2A ions (Alkaline Earth Metals) with the sulfate ion, SO_4^{2-} . By varying the amount of each group 2A ion present, you will determine the reactivity of each ion, and determine a pattern based on the position of the ion in the periodic table.

In this lab, we will observe the results of reactions involving alkaline earth metal ions and determine a pattern of reactivity of the elements in the family.

Materials:

- Goggles
- Test tubes
- Pipets
- Filter paper
- Test tubes
- Distilled water
- 0.10 M Mg^{2+}
- 0.10 M Ca^{2+}
- 0.10 M Sr^{2+}
- 0.10 M Ba^{2+}
- 0.10 M NO_3^-

Procedure:

1. Obtain and put on goggles.
2. Using a pipet, place ~10 drops of distilled water into wells A2-A5, B2-B5, C2-C5, and D2-D5. Wells A1, B1, C1, and D1 should be empty.
3. Using a different pipet for each ion solution, place ~20 drops of Mg^{2+} into well A2, ~20 drops of Ca^{2+} into well B2, ~20 drops of Sr^{2+} into well C2, and ~20 drops of Ba^{2+} into well D2.
4. Using the 10g pipet, remove half the solution from A2 and place it into well A1. Mix the solution in A2 with a test tube.
5. Using the 10g pipet, remove half the solution from A2 and place it into well A3. Mix the solution in A2 with the same test tube.
6. Using the 10g pipet, remove half the solution from A2 and place it into well A4. Mix the solution in A2 with the same test tube.
7. Using the 10g pipet, remove half the solution from A2 and place it into well A5. Mix the solution in A2 with the same test tube.

Periodicity Lab Answer

**United States. Environmental
Protection Agency**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the EPA text box.

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Table of Contents Periodicity Lab Answer

1. Understanding the eBook Periodicity Lab Answer
 - The Rise of Digital Reading Periodicity Lab Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodicity Lab Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodicity Lab Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodicity Lab Answer
 - Personalized Recommendations
 - Periodicity Lab Answer User Reviews and Ratings
 - Periodicity Lab Answer and Bestseller Lists
5. Accessing Periodicity Lab Answer Free and Paid eBooks
 - Periodicity Lab Answer Public Domain eBooks
 - Periodicity Lab Answer eBook Subscription Services

- Periodicity Lab Answer Budget-Friendly Options
- 6. Navigating Periodicity Lab Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodicity Lab Answer Compatibility with Devices
 - Periodicity Lab Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodicity Lab Answer
 - Highlighting and Note-Taking Periodicity Lab Answer
 - Interactive Elements Periodicity Lab Answer
- 8. Staying Engaged with Periodicity Lab Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodicity Lab Answer
- 9. Balancing eBooks and Physical Books Periodicity Lab Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodicity Lab Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodicity Lab Answer
 - Setting Reading Goals Periodicity Lab Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodicity Lab Answer
 - Fact-Checking eBook Content of Periodicity Lab Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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