

Percent Yield Practice

1. Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.



$$\frac{24.8 \text{ g CaCO}_3}{1} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \cdot \frac{1 \text{ mol CaO}}{1 \text{ mol CaCO}_3} \cdot \frac{56.08 \text{ g CaO}}{1 \text{ mol CaO}} = 13.9 \text{ g CaO} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{13.1}{13.9} \cdot 100 = \boxed{94.24\%} \Rightarrow \boxed{94.2\%}$$

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?



$$\frac{1.87 \text{ g Al}}{1} \cdot \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \cdot \frac{159.62 \text{ g CuSO}_4}{1 \text{ mol CuSO}_4} = 16.6 \text{ g CuSO}_4 \quad \text{Used}$$

$$\frac{9.65 \text{ g CuSO}_4}{159.62 \text{ g CuSO}_4} \cdot \frac{1 \text{ mol CuSO}_4}{1 \text{ mol CuSO}_4} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol CuSO}_4} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 3.84 \text{ g Cu}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{3.65}{3.84} \cdot 100 = \boxed{95.05\%} \quad \text{TY} \rightarrow \boxed{95.1\%}$$

3. If you start with 5.0 g of C_3H_8 and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?



$$\frac{5.0 \text{ g C}_3\text{H}_8}{44.11 \text{ g C}_3\text{H}_8} \cdot \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 8.2 \text{ g H}_2\text{O} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100$$

$$\frac{\text{PY} \cdot \text{TY}}{100} = \text{AY} = \frac{75 \cdot 8.2}{100} = \boxed{6.2 \text{ g H}_2\text{O}} \quad \text{AY}$$

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?



$$\frac{20 \text{ g CaO}}{1} \cdot \frac{1 \text{ mol CaO}}{56.08 \text{ g CaO}} \cdot \frac{1 \text{ mol Na}_2\text{O}}{1 \text{ mol CaO}} \cdot \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{20 \text{ g Na}_2\text{O}}$$

Percent Yield Practice Problems Answer Key

CO Houle



Percent Yield Practice Problems Answer Key:

Whispering the Secrets of Language: An Psychological Quest through **Percent Yield Practice Problems Answer Key**

In a digitally-driven world where monitors reign supreme and immediate connection drowns out the subtleties of language, the profound techniques and emotional nuances concealed within phrases often move unheard. However, set within the pages of **Percent Yield Practice Problems Answer Key** a interesting fictional prize pulsating with natural thoughts, lies a fantastic quest waiting to be undertaken. Composed by a skilled wordsmith, this marvelous opus invites viewers on an introspective trip, lightly unraveling the veiled truths and profound impact resonating within the fabric of each and every word. Within the emotional depths of this poignant evaluation, we shall embark upon a honest exploration of the book is key styles, dissect its charming publishing model, and succumb to the powerful resonance it evokes serious within the recesses of readers hearts.

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