

Elements of the Sea

Ionic equations

- If 2 solutions react to form a solid, a precipitation reaction occurs. Eg. $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{Na}^+(\text{aq}) + \text{Br}^-(\text{aq})$ goes to $\text{AgBr}(\text{s}) + \text{Na}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$

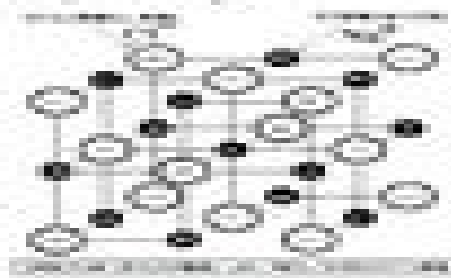
Then cancel spectators

$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq})$ goes to $\text{AgBr}(\text{s})$

- Always include state symbols
- Only split into ions if **aqueous and ionic**

Ionic solids

- Cations and anions are held together by strong electrostatic attractions in a **giant ionic lattice**.
- Each sodium ion is surrounded by 6 chloride ions **6:6 coordination**. It is a **simple cubic** structure.
- Ionic solids are hard and have a high melting point due to strong electrostatic attractions between ions.
- If ionic crystals contain water, the water molecules fit into the lattice the same way as regular ions. This is called the **water of crystallisation**. These crystals are said to be hydrated.
- If these hydrated crystals are heated, the water molecules are driven off and you are left with an **anhydrous solid**.



Ionic substances in solution

- As ionic substances dissolve in water, the ions become surrounded by water molecules and they spread throughout the solution. The hydrated ions are **randomly** arranged and behave **independently**.

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John Hadlett Harker, J. R. Backhurst



Salter's Elements of Sea Module Test:

Salter's Higher Chemistry, 1999 This work provides coverage of the content statements in the arrangements for Higher Chemistry organized by the three units in the course Energy Matters the World of Carbon and Chemical Reactions At the start of each unit students are given guidance on what they need to know and understand *The Marine Environment* J. M. A. Lenihan, William Whigham Fletcher, 1977 *The Chemical Engineer*, 1976 *Solar Energy Update*, 1978-04
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