

Part 2

GCSE Chemistry OCR A (Combined Science - **HIGHER**) Paper 2 Exam Walkthrough

Combined Science Higher - Chemistry

Walkthrough Mock Paper

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Look at the equation for the equilibrium reaction between sulfur dioxide, SO_2 , oxygen and sulfur trioxide, SO_3 , in a closed system.



(a) Predict the effect of adding more oxygen to the equilibrium mixture.

Explain your answer.

(1)

(b) Predict the effect of increasing the total pressure in the equilibrium mixture.

Explain your answer.

(1)

(c) Predict the effect of increasing the temperature of the equilibrium mixture.

Explain your answer.

(1)

(d) Sulfur dioxide for this reaction is made by burning sulfur.



Calculate the mass of sulfur needed to make 48 tonnes of sulfur dioxide.

Give your answer to 2 significant figures.

The relative atomic mass, A_r , of S is 32.1 and of O is 16.0.

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B Lingard



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