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**Rachel Sandford**



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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3\text{ mole H}_2\text{S} \times \frac{2\text{ mole H}_2\text{O}}{2\text{ mole H}_2\text{S}} \times 18\text{g H}_2\text{O} = 149.4\text{g H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   $160.0$  g  $\text{Al}$   $\times 1$  mol  $\text{Al}$   $\times 2$  mol  $\text{Al}_2\text{O}_3$   $\times 101.964$  g  $\text{Al}_2\text{O}_3$   $302.3$  g  $\text{Al}_2\text{O}_3$   $26.982$  g  $\text{Al}$   $4$  mol  $\text{Al}$   $1$  mol  $\text{Al}_2\text{O}_3$  c determine the percent yield yield actual  $\times 100$   $260.0$  g  $\times 100$   $86.01$  theoretical  $302.3$  g

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web 7 Na  $2$  C  $2$  O  $4$  is the limiting reactant percent yield  $86.68$  only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4  $33$  g of H  $2$  are left over 13

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