

### Quantum Numbers

- Used to \_\_\_\_\_ an \_\_\_\_\_ in an \_\_\_\_\_

#### $n$

- \_\_\_\_\_
- Represents \_\_\_\_\_ energy level of \_\_\_\_\_  
\_\_\_\_\_ of \_\_\_\_\_ in an \_\_\_\_\_

Example: What is the maximum number of electrons that can be in the \_\_\_\_\_ main energy level?

#### $l$

- The \_\_\_\_\_
- Describes the \_\_\_\_\_ within an \_\_\_\_\_
- \_\_\_\_\_ of orbital \_\_\_\_\_ possible is \_\_\_\_\_  
\_\_\_\_\_ = \_\_\_\_\_

### Orbital Shapes

designated \_\_\_\_\_

- level 1: \_\_\_\_\_
- level 2: \_\_\_\_\_
- level 3: \_\_\_\_\_
- level 4: \_\_\_\_\_

### How many electrons can each sublevel hold?

s = 1 orbital  $\times$  2 e<sup>-</sup>/orbital = \_\_\_\_\_e<sup>-</sup>  
 p = 3 orbitals  $\times$  2 e<sup>-</sup>/orbital = \_\_\_\_\_e<sup>-</sup>  
 d = 5 orbitals  $\times$  2 e<sup>-</sup>/orbital = \_\_\_\_\_e<sup>-</sup>  
 f = 7 orbitals  $\times$  2 e<sup>-</sup>/orbital = \_\_\_\_\_e<sup>-</sup>

# Note Taking Guide Episode 304 Key

**M Carnoy**



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