b) How many grams of phosphorus are required to react completely with 20.0 g of O1?

c) What is the theoretical yield in grams if you combine the amounts of reactants in part b?

A gaseous mixture containing 5.00 moles of H2 and 7.00 moles of Br2 reacts to form HBr.

a) Write a balanced chemical equation for this reaction.

b) Which reactant is limiting?

c) What is the theoretical yield for this reaction in moles?

d) What is the theoretical yield for this reaction in grams?

Sodium reacts with molecular hydrogen to form sodium hydride. A reaction mixture contains 10.0 g Na and 0.0235 g Hz.

a) Write a balanced chemical equation for this reaction.

c) What is the theoretical yield for this reaction in grams?

d) When this reaction is actually performed, 0.428 g of NaH is recovered. What is the percent yield of the reaction?

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**David Baud** 

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