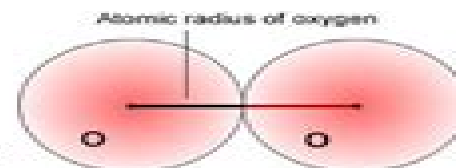


Vocabulary: Periodic Trends



Vocabulary

- **Atomic radius** – the distance from the nucleus to the outermost edge of the electron cloud.
 - The atomic radius can be determined by measuring the distance between two adjoining nuclei in a molecule comprised of two identical atoms, and then dividing by two.
 - For nonmetallic atoms that do not form diatomic molecules, the radius is determined by measuring the radius of atoms in covalent compounds.
 - For metallic atoms, the radius is determined by measuring the radius of atoms in metallic crystals.
 - The *Periodic Trends Gizmo* uses calculated atomic radii established by Enrico Clementi and others in 1967.
- **Electron affinity** – the change in energy that occurs when an electron is added to a neutral atom in the gaseous state, measured in kJ/mol.
 - Electron affinity refers to the tendency of an atom to attract electrons.
 - An element with a strong attraction for electrons will have a high negative electron affinity, since energy is released when an electron is added. The greater the magnitude of the negative electron affinity value, the stronger the attraction for electrons.
- **Electron cloud** – the region occupied by fast-moving electrons that surrounds the nucleus of an atom.
 - The electron cloud can be considered a probability distribution representing the space where electrons are most likely to be found.
- **Energy level** – a discrete region of space within the electron cloud where electrons are most likely to be found.
 - The number of principal energy levels within an atom is equal to that atom's period number on the periodic table.
 - There are a maximum of 7 principal energy levels within an atom.
- **Group** – a vertical column in the periodic table, numbered 1 to 18.
 - Groups are also called families.
 - Elements in the same group have the same number of valence electrons and similar chemical properties.



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George Graybill



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