

The Periodic Table

The Structure of the Periodic Table

- The periodic table displays the elements in order of increasing atomic number i.e. in order of increasing number of protons.
- Elements are arranged in 18 vertical columns called groups that are numbered 1 – 18.
- The seven horizontal rows in periodic table are called periods. The periods are numbered 1-7. The first period contains only two elements, hydrogen and helium.
- Elements within the same groups have similar chemical properties because they have the same number of valence electrons. The number of valence electrons is the same as the group number (in Roman numerals).

		GROUPS 1 - 18																					
PERIODS 1 - 7	1																			18			
	1	H 1.0079																			He 4.0026		
	2	Li 6.941	Be 9.0122															B 10.811	C 12.011	N 14.007	O 15.999	F 18.998	Ne 20.18
	3	Na 22.99	Mg 24.305															Al 26.982	Si 28.086	P 30.974	S 32.06	Cl 35.453	Ar 39.948
	4	K 39.098	Ca 40.078	Sc 44.956	Ti 47.88	V 50.942	Cr 51.996	Mn 54.938	Fe 55.847	Co 58.933	Ni 58.693	Cu 63.546	Zn 65.39	Ga 69.723	Ge 72.64	As 74.922	Se 78.96	Br 79.904	Kr 83.8				
	5	Rb 85.468	Sr 87.62	Y 88.906	Zr 91.224	Nb 92.906	Mo 95.94	Tc 97.907	Ru 101.07	Rh 102.91	Pd 106.42	Ag 107.87	Cd 112.41	In 114.82	Sn 118.71	Sb 121.76	Te 127.6	I 126.9	Xe 131.29				
	6	Cs 132.91	Ba 137.33	La 138.91	Hf 178.49	Ta 180.95	W 183.84	Re 186.21	Os 190.23	Ir 192.22	Pt 195.08	Au 196.97	Hg 200.59	Tl 204.38	Pb 207.2	Bi 208.98	Po 209	At 210	Rn 222				
	7	Fr 223	Ra 226	Ac 227	Rf 261	Ha 262	Sg 266	Nh 271	Hs 277	Mt 288	Ds 289	Ds 289	Ds 289	Ds 289	Ds 289	Ds 289	Ds 289	Ds 289	Ds 289				

Energy levels and period numbers

The electrons of an atom are arranged in energy levels. The period number tells us in which energy level an atom has its valence electrons.

For example:

Element	In which Group?	In which Period?	How many valence electrons	Which energy level are valence electrons?
Phosphorous	15	3	5	3 rd energy level
Potassium				
Calcium				
Oxygen				

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Ying-Ying Zheng



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