

Annotated periodic table of the elements

1	H 1.01 Hydrogen												13	14	15	16	17	18
2	3	4	Main common oxidation states: electron affinity (kJ/mol): (blue: positive)										31	32	33	34	35	36
3	Li 6.94 Lithium	Be 9.01 Beryllium											51	52	53	54	55	56
4	Na 22.99 Sodium	Mg 24.31 Magnesium											71	72	73	74	75	76
5	K 39.10 Potassium	Ca 40.08 Calcium											91	92	93	94	95	96
6	Rb 85.47 Rubidium	Sr 87.62 Strontium											111	112	113	114	115	116
7	Cs 132.91 Cesium	Ba 137.33 Barium											131	132	133	134	135	136
8	Fr 223.02 Francium	Ra 226.03 Radium											151	152	153	154	155	156
9													171	172	173	174	175	176
10													191	192	193	194	195	196
11													211	212	213	214	215	216
12													231	232	233	234	235	236
13													251	252	253	254	255	256
14													271	272	273	274	275	276
15													291	292	293	294	295	296
16													311	312	313	314	315	316
17													331	332	333	334	335	336
18													351	352	353	354	355	356
19													371	372	373	374	375	376
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21													411	412	413	414	415	416
22													431	432	433	434	435	436
23													451	452	453	454	455	456
24													471	472	473	474	475	476
25													491	492	493	494	495	496
26													511	512	513	514	515	516
27													531	532	533	534	535	536
28													551	552	553	554	555	556
29													571	572	573	574	575	576
30													591	592	593	594	595	596
31													611	612	613	614	615	616
32													631	632	633	634	635	636
33													651	652	653	654	655	656
34													671	672	673	674	675	676
35													691	692	693	694	695	696
36													711	712	713	714	715	716
37													731	732	733	734	735	736
38													751	752	753	754	755	756
39													771	772	773	774	775	776
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41													811	812	813	814	815	816
42													831	832	833	834	835	836
43													851	852	853	854	855	856
44													871	872	873	874	875	876
45													891	892	893	894	895	896
46													911	912	913	914	915	916
47													931	932	933	934	935	936
48													951	952	953	954	955	956
49													971	972	973	974	975	976
50													991	992	993	994	995	996
51													1011	1012	1013	1014	1015	1016
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59													1171	1172	1173	1174	1175	1176
60													1191	1192	1193	1194	1195	1196
61													1211	1212	1213	1214	1215	1216
62													1231	1232	1233	1234	1235	1236
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154													3071	3072	3073	3074	3075	3076
155																		

- Group names*
- 1 Hydrogen & the alkali metals
 - 2 Alkaline earth metals
 - 11 Coinage metals
 - 12 Volatile metals
 - 13 Lanthanoids
 - 14 Chalcogens
 - 15 Pnictogens
 - 16 Chalcogens
 - 17 Halogens
 - 18 Helium group
- *Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

87 La 138.91 Lanthanum	88 Ce 140.12 Cerium	89 Pr 140.91 Praseodymium	90 Nd 144.24 Neodymium	91 Pm 144.91 Promethium	92 Sm 150.36 Samarium	93 Eu 151.96 Europium	94 Gd 157.25 Gadolinium	95 Tb 158.93 Terbium	96 Dy 162.50 Dysprosium	97 Ho 164.93 Holmium	98 Er 167.26 Erbium	99 Tm 168.93 Thulium	100 Yb 173.05 Ytterbium
101 Ac 227.03 Actinium	102 Th 232.04 Thorium	103 Pa 231.04 Protactinium	104 U 238.03 Uranium	105 Np 237.05 Neptunium	106 Pu 244.06 Plutonium	107 Am 243.06 Americium	108 Cm 247.07 Curium	109 Bk 247.07 Berkelium	110 Cf 251.08 Californium	111 Es 252.08 Einsteinium	112 Fm 257.10 Fermium	113 Md 258.10 Mendelevium	114 No 259.10 Nobelium

Radioactive

Oganesson is likely to be either a metalloid or a post-transition metal

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zirconates, titanates, and tantalates.

Metals

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

* as liquid P, the most stable form in ambient conditions.

Group 1
The alkali metals Li, Na, K, Rb, Cs, and Fr are all highly reactive. The alkali metals are found in the Earth's crust as salts. The alkali metals are found in the Earth's crust as salts. The alkali metals are found in the Earth's crust as salts.

Nonmetals

Covalent nonmetals
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically covalent nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePtF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

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L Towner

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